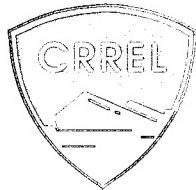


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SPECIAL REPORT



FREZCHEM: A Chemical-Thermodynamic Model for Aqueous Solutions at Subzero Temperatures

Giles M. Marion and Steven A. Grant

July 1994



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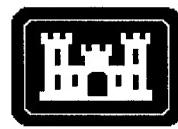
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Abstract

This report documents a FORTRAN version of the Spencer-Møller-Weare chemical thermodynamic model for aqueous electrolyte solutions at subzero temperatures (FREZCHEM). FREZCHEM is structured to predict the chemical composition and unfrozen water of aqueous solutions between -60 °C and +25 °C at atmospheric pressure (0.101325 MPa). FREZCHEM includes two reaction pathways: 1) freezing at variable temperature and fixed total water and 2) evaporation at variable water and fixed temperature. Activity coefficients and the activity of water are calculated using the Pitzer equations, which are valid to high solution ionic strengths ($\approx 20 \text{ mol kg}^{-1}$). Fifteen chloride and sulfate salts of sodium, potassium, calcium, and magnesium are included in the model. Predicted and experimental measurements of solute molalities and the unfrozen water fraction during seawater freezing are in good agreement. At -50 °C, 0.3% of seawater remains unfrozen with 99.7% of Na and 95.5 % of Cl having precipitated into one of four salts. FREZCHEM should find many applications in physicochemical studies of aqueous solutions and freezing.

For conversion of SI metric units to U.S./British customary units of measurement consult ASTM Standard E380-89a, *Standard Practice for Use of the International System of Units*, published by the American Society for Testing and Materials, 1916 Race St., Philadelphia, Pa. 19103.



**U.S. Army Corps
of Engineers**
Cold Regions Research &
Engineering Laboratory

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Giles M. Marion and Steven A. Grant

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PREFACE

This report was prepared by Dr. Giles M. Marion and Dr. Steven A. Grant, Research Physical Scientists, of the Geochemical Sciences Branch, Research Division, U. S. Army Cold Regions Research and Engineering Laboratory. Funding was provided by DA Project 4A161102AT24, *Research in Snow, Ice and Frozen Ground*, Tasks EC and SC, Work Units B03, *Soil Solute Interactions at Low Temperature*, and F02, *Chemical Processes in Frozen Soil*, and by Strategic Environmental Research and Development Program (SERDP) Work Unit, *Fate and Transport in Seasonally Frozen Soil and Discontinuous Permafrost*.

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FREZCHEM: A Chemical-Thermodynamic Model for Aqueous Solutions at Subzero Temperatures

GILES M. MARION AND STEVEN A. GRANT

INTRODUCTION

The presence of solutes lowers the freezing-point of aqueous solutions; and when solutions freeze, ice formation largely excludes solutes, concentrating them in the unfrozen brine which, in turn, further lowers the brine freezing point. For example, when the temperature of a 1.0 molal NaCl solution is lowered below 0 °C, ice starts to form at approximately -3 °C (Fig. 1). As ice forms, the NaCl molality increases and the temperature drops along the ice-solution line, assuming that equilibrium is maintained, until the eutectic is reached at -21.3 °C and 5.17 molal (Fig. 1). At this point, the residual solution freezes, forming a mixture of ice and hydrohalite [NaCl•2H₂O(cr)]. Clearly, complex interactions exist between solute concentrations and the freezing process.

A number of chemical thermodynamic models for aqueous electrolyte solutions have been developed in recent years (Sposito and Mattigod 1979, Plummer et al. 1988, Spencer et al. 1990). However, only the Spencer-Møller-Weare model deals explicitly with aqueous solutions at subzero temperatures. This model accounts for the precipitation of the chloride and sulfate salts of sodium, potassium, calcium, and magnesium over the temperature range from -60 °C to +25 °C at atmospheric pressure (0.101325 MPa). While the parameters used in the model have been published (Spencer et al. 1990), a working version of the model has not been published. Such a working model would be extremely useful for geochemists and geophysicists interested in the complex interactions between solutes and the freezing process. This preliminary report will document a FORTRAN version of the Spencer-Møller-Weare model, which we have chosen to call FREZCHEM.

MODEL STRUCTURE

FREZCHEM is a chemical thermodynamic "equilibrium" model. This model will calculate the equilibrium composition of aqueous solutions at specified temperatures, but will not provide any information on the time required to reach, or the path of change to, the equilibrium state. FREZCHEM calculates equilibrium between water and ice and between dissolved and solid phase salts.

For water and ice at equilibrium,

$$\begin{aligned}\mu_i = \mu_i^\circ + RT \ln(a_{H_2O(cr,I)}) &= \mu_w = \mu_w^\circ + \\ RT \ln(a_{H_2O(l)}) &\end{aligned}\quad (1)$$

where R is the gas constant, T is absolute temperature, μ_i and μ_w are the chemical potentials of ice and water, μ_i° and μ_w° are the standard chemical potentials, and $a_{H_2O(cr,I)}$ and $a_{H_2O(l)}$ are the activities of water in the form of ice and liquid water. Equation 1 can be rearranged to yield

$$\frac{\mu_i^\circ - \mu_w^\circ}{RT} = \ln \left(\frac{a_{H_2O(l)}}{a_{H_2O(cr,I)}} \right) = \frac{\Delta_{fus}G^\circ}{RT} \quad (2)$$

where $\Delta_{fus}G^\circ$ is the standard Gibbs energy of fusion ($\mu_i^\circ - \mu_w^\circ$). At equilibrium with respect to pure ice, $a_{H_2O(cr,I)} = 1.0$. Therefore

$$(a_{H_2O(l)}) = \exp[\Delta_{fus}G^\circ/RT] = K \quad (3)$$

the activity of water is equal to the equilibrium constant (K) at a given temperature. For equilibrium with respect to a salt (e.g., hydrohalite)

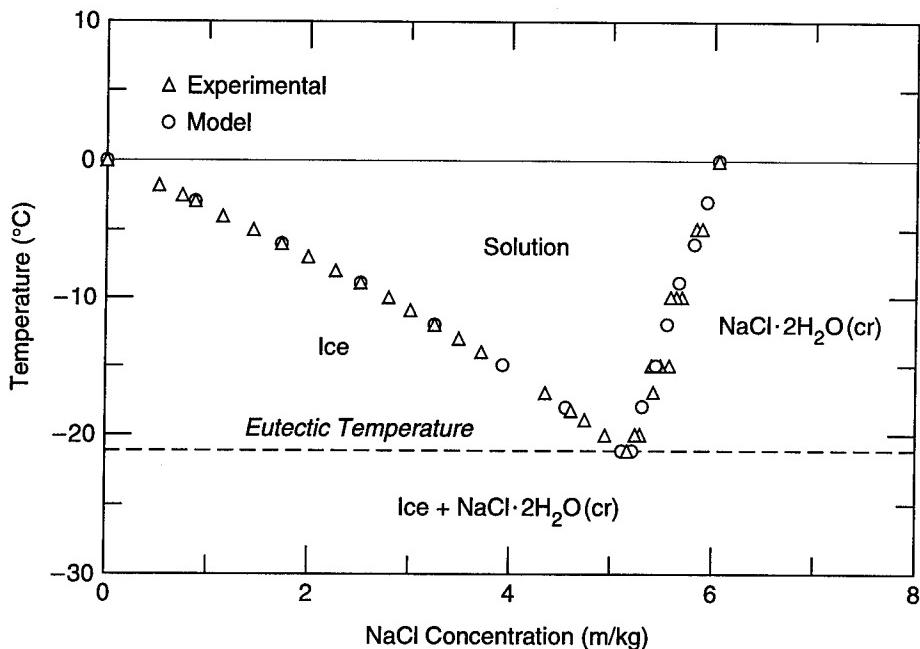
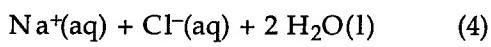
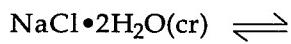


Figure 1. A stability diagram for the $\text{NaCl}-\text{H}_2\text{O}$ system at subzero temperatures. Experimental measurements are from Bukshtein et al. (1953) and Hall et al. (1988).



and

$$(a_{\text{Na}^+(\text{aq})})(a_{\text{Cl}^-(\text{aq})})(a_{\text{H}_2\text{O(l)}})^2 = K \quad (5)$$

assuming that the activity of the pure salt $[\text{NaCl} \cdot 2\text{H}_2\text{O(cr)}] = 1.0$. The relation between activity (a) and molality (m) of a substance "B" is given by

$$a_B = \gamma_B m_B \quad (6)$$

where γ is the activity coefficient. If molalities are given, to calculate activities requires an extra-thermodynamic model for activity coefficients.

The Pitzer equations are used to calculate the activity of water and the activity coefficients for soluble ions. These equations allow calculation of activity coefficients to high ionic strengths in complex solutions (Plummer et al. 1988, Spencer et al. 1990, Pitzer 1991).

The osmotic coefficient (ϕ), and single-ion activity coefficients for cations (γ_M) and anions (γ_X) are given by

$$\begin{aligned} (\phi - 1) &= 2 / (\sum m_i) \left\{ -A \Phi I^{1.5} / (1 + b I^{0.5}) + \right. \\ &\quad \left. \Sigma \Sigma m_c m_a (B \Phi_{ca} + Z C_{ca}) \right\} \\ &+ \Sigma \Sigma m_c m_c (\Phi^\phi_{cc'} + \Sigma m_a \Psi_{ca'c}) + \\ &\quad \left. \Sigma \Sigma m_a m_a (\Phi^\phi_{aa'} + \Sigma m_c \Psi_{aa'c}) \right\} \end{aligned} \quad (7)$$

$$\begin{aligned} \ln(\gamma_M) &= z_M^{-2} F + \Sigma m_a (2B_{Ma} + Z C_{Ma}) + \\ &\quad \Sigma m_c (2\Phi_{Mc} + \Sigma m_a \Psi_{Mca}) \\ &+ \Sigma \Sigma m_a m_a \Psi_{aaM} + |z_M| \Sigma \Sigma m_c m_a C_{ca} \end{aligned} \quad (8)$$

$$\begin{aligned} \ln(\gamma_X) &= z_X^{-2} F + \Sigma m_c (2B_{cx} + Z C_{cx}) + \\ &\quad \Sigma m_a (2\Phi_{Xa} + \Sigma m_c \Psi_{Xac}) \\ &+ \Sigma \Sigma m_c m_c \Psi_{ccX} + |z_X| \Sigma \Sigma m_c m_a C_{ca} \end{aligned} \quad (9)$$

where $A\Phi$ is the Debye-Hückel constant, b is a constant (=1.2), B^ϕ , C , Φ^ϕ , and Ψ are salt interaction coefficients, z_i is the charge number of the i th ion, I is the ionic strength defined by

$$I = 0.5 \sum m_i z_i^2, \quad (10)$$

Z is defined by

Table 1. Minerals currently in the FREZCHEM model.

1. $\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}(\text{cr})$ (mirabilite)
2. $\text{Na}_2\text{SO}_4(\text{cr})$ (thenardite)
3. $\text{K}_2\text{SO}_4(\text{cr})$ (arcanite)
4. $\text{MgSO}_4 \cdot 6\text{H}_2\text{O}(\text{cr})$ (hexahydrite)
5. $\text{MgSO}_4 \cdot 7\text{H}_2\text{O}(\text{cr})$ (epsomite)
6. $\text{MgSO}_4 \cdot \text{K}_2\text{SO}_4 \cdot 6\text{H}_2\text{O}(\text{cr})$ (picromerite)
7. $\text{NaCl}(\text{cr})$ (halite)
8. $\text{NaCl} \cdot 2\text{H}_2\text{O}(\text{cr})$ (hydrohalite)
9. $\text{KCl}(\text{cr})$ (sylvite)
10. $\text{CaCl}_2 \cdot 6\text{H}_2\text{O}(\text{cr})$ (antarcticite)
11. $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}(\text{cr})$ (bischofite)
12. $\text{MgCl}_2 \cdot 8\text{H}_2\text{O}(\text{cr})$
13. $\text{MgCl}_2 \cdot 12\text{H}_2\text{O}(\text{cr})$
14. $\text{KMgCl}_3 \cdot 6\text{H}_2\text{O}(\text{cr})$ (carnallite)
15. $\text{CaCl}_2 \cdot 2\text{MgCl}_2 \cdot 12\text{H}_2\text{O}(\text{cr})$ (tachyhydrite)

$$Z = \sum m_i |z_i| \quad (11)$$

and F is a complex function of A^ϕ , I , and the salt interaction coefficients (Plummer et al. 1988, Pitzer 1991). Given ϕ from eq 7, the activity of water is

defined by

$$a_{\text{H}_2\text{O}(\text{l})} = \exp(-\phi \sum m_i / 55.50837). \quad (12)$$

The equilibrium constants and the Pitzer equation parameters for the temperature range -60°C to $+25^\circ\text{C}$ were published by Spencer et al. (1990) as functions of temperature using equations of the form

$$P(T) = a_1 + a_2 T + a_6 T^2 + a_9 T^3 + a_3/T + a_4 \ln(T) \quad (13)$$

where $P(T)$ is the Pitzer equation parameter directly or the function, $[-\Delta_f G^\circ / RT]$ (where $\Delta_f G^\circ$ is the standard reaction Gibbs energy), for equilibrium constants [$K = \exp(-\Delta_f G^\circ / RT)$] and T is thermodynamic temperature (K). Refer to Spencer et al. (1990) for the specific parameters used in the model (their Tables 1, 2, and 3). The only additional parameter, beyond those used in the Spencer-Møller-Weare model, is $\Psi(\text{Ca}, \text{Mg}, \text{SO}_4) = 0.024$, which is treated as a constant independent of temperature (Pitzer 1991).

Table 2. A listing of chemical species and their numerical designation in the FREZCHEM model.

A. Solution and atmospheric species

#	Species	#	Species	#	Species
1	$\text{Na}^+(\text{aq})$	11	$\text{Cl}^-(\text{aq})$	21	$\text{CO}_2(\text{aq})$
2	$\text{K}^+(\text{aq})$	12	$\text{SO}_4^{2-}(\text{aq})$	22	$\text{CaSO}_4^\circ(\text{aq})$
3	$\text{Ca}^{2+}(\text{aq})$	13	$\text{OH}^-(\text{aq})$	23	$\text{MgSO}_4^\circ(\text{aq})$
4	$\text{Mg}^{2+}(\text{aq})$	14	$\text{HCO}_3^-(\text{aq})$	24	
5	$\text{H}^+(\text{aq})$	15	$\text{CO}_3^{2-}(\text{aq})$	25	
6		16		26	
7		17		27	
8		18		28	
9		19		29	$\text{CO}_2(\text{g})$
10		20		30	$\text{H}_2\text{O}(\text{l})$

B. Solid phase species

#	Species	#	Species	#	Species
31	$\text{H}_2\text{O}(\text{cr}, \text{l})$	41	$\text{Na}_2\text{SO}_4 \cdot 10\text{H}_2\text{O}(\text{cr})$	51	
32	$\text{NaCl} \cdot 2\text{H}_2\text{O}(\text{cr})$	42	$\text{Na}_2\text{SO}_4(\text{cr})$	52	
33	$\text{NaCl}(\text{cr})$	43	$\text{MgSO}_4 \cdot 6\text{H}_2\text{O}(\text{cr})$	53	
34	$\text{KCl}(\text{cr})$	44	$\text{MgSO}_4 \cdot 7\text{H}_2\text{O}(\text{cr})$	54	
35	$\text{CaCl}_2 \cdot 6\text{H}_2\text{O}(\text{cr})$	45	$\text{K}_2\text{SO}_4(\text{cr})$	55	
36	$\text{MgCl}_2 \cdot 6\text{H}_2\text{O}(\text{cr})$	46	$\text{MgSO}_4 \cdot \text{K}_2\text{SO}_4 \cdot 6\text{H}_2\text{O}(\text{cr})$	56	
37	$\text{MgCl}_2 \cdot 8\text{H}_2\text{O}(\text{cr})$	47		57	
38	$\text{MgCl}_2 \cdot 12\text{H}_2\text{O}(\text{cr})$	48		58	
39	$\text{KMgCl}_3 \cdot 6\text{H}_2\text{O}(\text{cr})$	49		59	
40	$\text{CaCl}_2 \cdot 2\text{MgCl}_2 \cdot 12\text{H}_2\text{O}(\text{cr})$	50		60	

Table 3. FREZCHEM model output for freezing of seawater at -30 °C.

Temp (K) 243.15	Ion. Str. 7.8770	AH ₂ O 0.74696	Phi 1.6390	H ₂ O (g) 33.120	Ice (g) 948.20
<i>Solution Species</i>	<i>Initial Conc.</i>	<i>Final Conc.</i>	<i>Act. Coef.</i>	<i>Activity</i>	<i>Moles</i>
Na	0.48695	1.6940	0.54243	0.91885	0.56104E-01
K	0.10630E-01	0.32095	0.19599	0.62905E-01	0.10630E-01
Ca	0.95300E-02	0.28673	0.51732	0.14833	0.94967E-02
Mg	0.55160E-01	1.6654	1.1282	1.8790	0.55160E-01
Cl	0.56818	5.9081	1.8095	10.691	0.19568
SO ₄	0.29390E-01	0.55838E-02	0.32328	0.18051E-02	0.18494E-03
CaSO ₄	0.00000	0.10058E-02	1.0000	0.10058E-02	0.33313E-04
MgSO ₄	0.00000	0.42492E-05	1.0000	0.42492E-05	0.14074E-06
H ₂ O(l)				0.74696	1.8384
<i>Solid Species</i>	<i>Moles</i>	<i>Equil. Constant</i>			
Ice	52.633	0.74693			
NaCl ₂ •2H ₂ O	0.37250	5.4770			
NaCl	0.00000	19.127			
KCl	0.00000	1.0399			
CaCl ₂ •6H ₂ O	0.00000	695.12			
MgCl ₂ •6H ₂ O	0.00000	40585.0			
MgCl ₂ •8H ₂ O	0.00000	1403.3			
MgCl ₂ •12H ₂ O	0.00000	27.775			
KMgCl ₃ •6H ₂ O	0.00000	1054.2			
CaCl ₂ •2MgCl ₂ •12H ₂ O	0.00000	0.17775E+21			
Na ₂ SO ₄ •10H ₂ O	0.29172E-01	0.81213E-04			
Na ₂ SO ₄	0.00000	0.73236			
MgSO ₄ •6H ₂ O	0.00000	3.1211			
MgSO ₄ •7H ₂ O	0.00000	0.93967E-01			
K ₂ SO ₄	0.00000	0.61722E-02			
MgSO ₄ •K ₂ SO ₄ •6H ₂ O	0.00000	0.21935E-04			
CaSO ₄ (ion-pair)		0.26618			
MgSO ₄ (ion-pair)		798.18			
Iterations = 39					

The specific salts included in the present model are listed in Table 1. In addition to these salts, the model also explicitly considers ion-pairs for Ca-SO₄ and MgSO₄.

MATHEMATICAL ALGORITHM

Two basic mathematical approaches are used in chemical thermodynamic models to calculate equilibrium concentrations and activities: Gibbs energy minimization or solution of sets of nonlinear equations. The Spencer et al. (1990) model explicitly uses Gibbs energy minimization. The

PHRQPITZ model (Plummer et al. 1988) uses the nonlinear equation solution method. Both approaches ultimately provide the same answers.

The FREZCHEM model solves sets of nonlinear equations that include ice–water and mineral equilibria. These equations may, in principle, be solved either simultaneously or sequentially. FREZCHEM uses the sequential approach, which treats each reaction as an independent reaction (Fig. 2). The solution of the set of nonlinear equations occurs by iterating through the sequence many times. Advantages of the sequential approach vis-a-vis the simultaneous or Gibbs energy minimization approaches include 1) programming

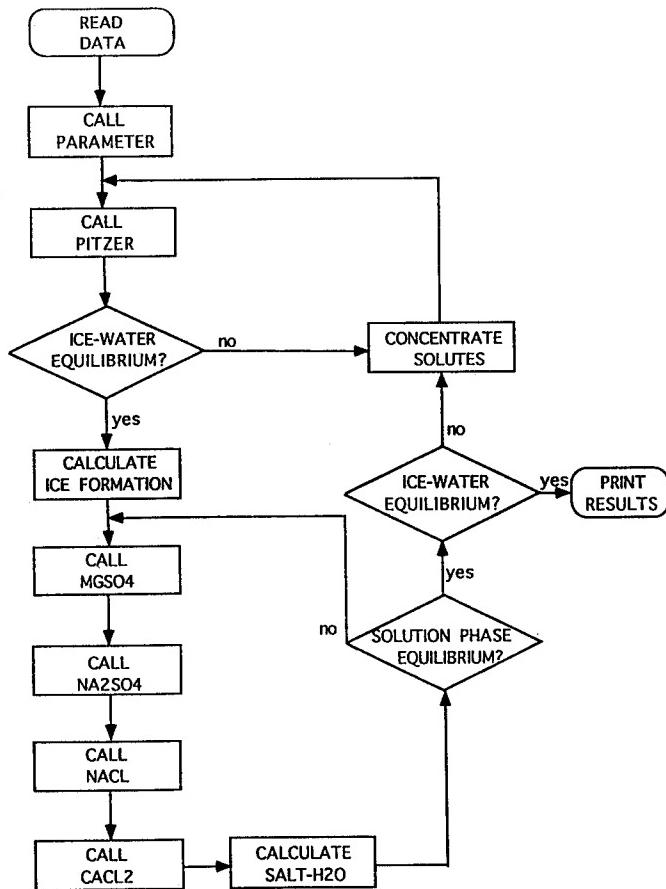


Figure 2. The FREZCHEM model flowchart.

simplicity since only one reaction is considered in separate subroutines, 2) ease of adding or subtracting new reactions since they can be inserted or deleted at any point in the sequence, and 3) the ease of program operation for both simple and complex aqueous solutions. With respect to the last point, only subroutines containing the elements in the "input" listing are used. For example, if only Na and Cl are present, then the program will only "call" the subroutine dealing with NaCl equilibrium and will bypass the remainder. For complex aqueous solutions such as seawater, the program will "call" every subroutine. The major disadvantage of the sequential approach is that it can be slow to converge. Another problem with the current model is convergence. At times, the program diverges or oscillates and never arrives at the equilibrium concentrations.

The program assumes that the initial aqueous solution is unfrozen and calculates the amount of ice, if any, that will form at the specified salt con-

centrations and temperature. The program first tests to see if ice and water are at equilibrium at the specified temperature. The criteria are

$$\text{IF } \alpha_{\text{H}_2\text{O(l)}} > K \text{ THEN freeze} \quad (14)$$

$$\text{IF } \alpha_{\text{H}_2\text{O(l)}} = K \text{ THEN equilibrium} \quad (15)$$

$$\text{IF } \alpha_{\text{H}_2\text{O(l)}} < K \text{ THEN water is stable phase} \quad (16)$$

where $\alpha_{\text{H}_2\text{O(l)}}$ is the activity of water calculated from the Pitzer equations (eq 7 and 12) and K is the equilibrium constant for ice-water (eq 3). If $\alpha_{\text{H}_2\text{O(l)}} > K$, this implies that water is thermodynamically unstable and ice (stable phase) should form. Ice formation is calculated indirectly. First, solute concentrations are increased proportionally, which decreases $\alpha_{\text{H}_2\text{O(l)}}$ (eq 12) until $\alpha_{\text{H}_2\text{O(l)}} = K$ and equilibrium is reached. Ice formation is calculated by

$$\text{Ice} = \text{TotalH}_2\text{O} - \text{Water}_f \quad (17)$$

$$\text{Water}_f = \text{Water}_i (\Sigma m_i / \Sigma m_f) \quad (18)$$

where Water_f, Water_i, and TotalH₂O are the final water, initial water, and total H₂O (ice and water), m_i are the initial molalities of solutes, and m_f are the final molalities. These calculations assume that the ice phase is a pure phase. If $a_{\text{H}_2\text{O}(l)} < K$, this implies that water is the thermodynamically stable phase and ice should not form at the specified salt concentrations and temperature.

In addition to freezing, FREZCHEM will also evaporate solutions. In contrast to freezing, the evaporation of water is not controlled by an equilibrium relation. The evaporation routine simply removes water incrementally from the solution, at a fixed temperature, according to the user's specifications. The program will not allow evaporation if ice is thermodynamically stable at the specified temperature and solution concentrations.

After ice and water are equilibrated, the program cycles through subroutines to handle mineral equilibria (Fig. 2). At present, there are 11 subroutines to handle the 15 mineral phases (Table 1) and the two ion-pair equilibria. Minerals that are identical in chemical composition except for waters of hydration are handled in the same subroutine [e.g., NaCl(cr) and NaCl·2H₂O(cr)].

After cycling through the mineral equilibrium equations, the program calculates the amount of water associated with hydrated salts (Fig. 2). Then the program tests for overall mineral equilibrium. In FREZCHEM, the initial and final solution phase concentrations must agree to within $\pm 0.1\%$ for all solution phase cations and anions. In general if more than one mineral phase is precipitating, the first cycle through the mineral subroutines will not suffice to establish overall mineral equilibrium and the equations must be solved repeatedly.

Eventually mineral equilibrium is established. Then the program tests to see if ice and water are still in equilibrium or $a_{\text{H}_2\text{O}(l)} < K$. For equilibrium, $a_{\text{H}_2\text{O}(l)}$ must agree with K to within ± 0.00005 . If ice and water are not in equilibrium and $a_{\text{H}_2\text{O}(l)} > K$, then the program cycles through the ice-water equilibrium and subsequently mineral equilibrium subroutines. Eventually the program converges to within the numerical equilibrium criteria established in the program, and output is sent to the print files.

FREZCHEM PROGRAM

A listing of the FREZCHEM FORTRAN program is in Appendix A. A major effort was made to adhere to the FORTRAN 77 ANSI Standard in writing this program. For example, FREZCHEM uses an 80-column statement line where columns 1–5 are statement labels, column 6 is a continuation column, columns 7–72 are used for statements, and columns 73–80 are used for commentary. Variable names are ≤ 6 characters. FREZCHEM is a FORTRAN translation of an earlier program written in TrueBASIC. FORTRAN assumes, unless otherwise declared, that variables beginning with the letters "IJKLMN" are integer variables and all other initial letters refer to real variables. In order to retain the same variable names between programs, this restriction required formal declaration of some real variables in the FORTRAN version. FREZCHEM is currently running on a Macintosh Quadra 800 computer using the MacFortran II compiler (version 3.2, Absoft 1991).

The FREZCHEM program consists of a main program called *FREZCHEM* and 15 subroutines. *FREZCHEM* reads in data from screen queries, initializes some parameters, freezes or evaporates the aqueous solution, calculates the water of hydration, tests for chemical thermodynamic equilibria, and most importantly "calls" the various subroutines (Fig. 2).

Subroutine *PARAMETER* calculates the parameters for the Pitzer equations and the equilibrium constants as functions of temperature (eq 13). For the most part, these parameters are taken from Spencer et al. (1990). Subroutine *PITZER* calculates activity coefficients and the activity of water using the Pitzer and other equations (eq 7–12). Subroutine *INTERACTION* calculates the higher-order electrostatic interaction terms for the Pitzer equations. Subroutine *PPRINT* prints the results. The remaining 11 subroutines deal with mineral equilibria. For example, Subroutine *MGCL2* determines which MgCl₂ salt is thermodynamically stable and equilibrates, if necessary, the salt with the solution phase. Other mineral subroutines perform similar functions for specific minerals.

FREZCHEM was structured to be user-expandable. To alter the model requires knowing how chemical species are explicitly designated in the model. Table 2 summarizes the chemical species currently in or anticipated to be included in the

model in the near future. Species 1–10 are reserved for soluble cations; species 11–20 are reserved for soluble anions; species 21–30 are reserved for neutral species, both solution phase and gas phase. The solid phase species are designated with numbers 31–60. Blank spaces in Table 2 are available for future additions.

PROGRAM INPUT AND OUTPUT

Input to FREZCHEM is through screen queries that request the molal concentrations for Na, K, Ca, Mg, Cl, and SO₄. FREZCHEM is structured to cycle through a sequence of temperature changes (freezing) or water removals (evaporation). For a temperature change, the screen queries request the initial temperature, the final temperature, and the incremental temperature change. For example, to calculate the freezing of seawater between 273.15 K (0 °C) and 223.15 K (-50 °C) by 5 K intervals, enter: 273.15, 223.15, and 5.0. Because the model assumes an initially unfrozen solution, enter the temperature sequence from high (initial) to low (final). For evaporative removal, the screen queries request the temperature (assumed fixed for evaporation), the final total water content (grams), and the incremental water removal (grams). To evaporate water from 1000 g (the fixed initial water content for both freezing and evaporation) to 100 g by 50-g increments at 0.0 °C, enter 273.15, 100, and 50 for the evaporative screen queries.

Output from the program is to a file called "FrData" which can be sent to a printer. A freezing simulation for seawater at 243.15 K (-30 °C) is given in Table 3, and an evaporative (1000 g to 100 g) simulation for seawater at 0 °C is given in Table 4. After the title, the first line of output gives the temperature, ionic strength, activity of water (AH₂O), osmotic coefficient (phi), and the final calculated amounts of water and ice. The basic unit of mass in the model is 1.0 kg water. For the seawater simulation at -30 °C, the final distribution of water is 948.20 g of ice, 33.12 g water, 13.42 g of water in hydrohalite (0.3725 moles), and 5.26 g of water in mirabilite (0.0292 moles), which adds up to exactly 1000.00 g (Table 3). The final distribution of water in the evaporative scenario at 0 °C is 96.93 g water and 3.07 g water in mirabilite (0.017043 moles), which adds up to exactly 100.00 g (Table 4).

The solution phase species are grouped together in the output. This output includes initial concen-

tration (the program input), final concentration, activity coefficient, activity, and the moles remaining in the solution phase. Because the basic unit of mass is 1.0 kg water, the initial concentration [moles/kg (water)] is also equal to the total moles. A difference between "Initial Conc." and "Moles" in the solution phase output is attributable to either salt precipitation or ion-pair formation. For example, the initial concentration (total moles) of Na is 0.48695. The distribution of Na at -30 °C is 0.05610 moles in the solution, 0.37250 moles as hydrohalite, and 0.05834 moles as mirabilite, which adds up to 0.48694 moles of Na (Table 3). For the seawater evaporative simulation, the final distribution of Na is 0.45286 moles in the solution and 0.03409 moles as mirabilite, which adds up to 0.48695 moles of Na (Table 4)

The solid phase species are also grouped together in the output. This output includes the moles of solid phase species and the equilibrium constant for the reaction. For the seawater freezing simulation, ice, hydrohalite, and mirabilite have precipitated at -30 °C (Table 3). For the seawater evaporative simulation, only mirabilite appears to have precipitated at 0 °C (Table 4); however, see *Program Limitations* (below) for a discussion of probable gypsum precipitation. Whenever ice is present, the activity of water (0.74696) must agree with the equilibrium constant for ice-water (0.74693) (Table 3) to within ±0.00005 units, which is the convergence criterion used in the model. For salts that are precipitating, the appropriate solution phase activity product should equal the equilibrium constant. For NaCl•2H₂O(cr), the activity product is equal to (0.91885) × (10.691) × (0.74696)² = 5.48 which agrees with the equilibrium constant of 5.48 (Table 3).

The final output are the equilibrium constants for the CaSO₄ and MgSO₄ ion-pairs and the number of iterations through the equilibrium routines before convergence.

PROGRAM LIMITATIONS

FREZCHEM has convergence problems especially at high ionic strengths (>15 molal), where activity coefficients are rapidly changing, and at junctions, where new phases begin to precipitate. The program contains a routine to catch mathematical problems or special circumstances. If "iterations" reaches 300, a message to this effect is

Table 4. FREZCHEM model output for evaporation of seawater at 0 °C.

Temp (K) 273.15	Ion. Str. 6.8871	AH ₂ O 0.75383	Phi 1.3728	H ₂ O (g) 96.930	Ice (g) 0.00000
<i>Solution Species</i>	<i>Initial Conc.</i>	<i>Final Conc.</i>	<i>Act. Coef.</i>	<i>Activity</i>	<i>Moles</i>
Na	0.48695	4.6721	0.81929	3.8278	0.45286
K	0.10630E-01	0.10967	0.39751	0.43593E-01	0.10630E-01
Ca	0.95300E-02	0.92284E-01	1.0610	0.97912E-01	0.89451E-02
Mg	0.55160E-01	0.56905	2.1926	1.2477	0.55158E-01
Cl	0.56818	5.8618	1.2301	7.2105	0.56818
SO ₄	0.29390E-01	0.12133	0.30951E-01	0.37552E-02	0.11760E-01
CaSO ₄	0.00000	0.60346E-02	1.0000	0.60346E-02	0.58493E-03
MgSO ₄	0.00000	0.19557E-04	1.0000	0.19557E-04	0.18957E-05
H ₂ O(l)				0.75383	5.3804
<i>Solid Species</i>	<i>Moles</i>	<i>Equil. Constant</i>			
Ice	0.00000	1.0002			
NaCl•2H ₂ O	0.00000	17.957			
NaCl	0.00000	31.339			
KCl	0.00000	3.7731			
CaCl ₂ •6H ₂ O	0.00000	1862.7			
MgCl ₂ •6H ₂ O	0.00000	55958.			
MgCl ₂ •8H ₂ O	0.00000	7080.2			
MgCl ₂ •12H ₂ O	0.00000	708.26			
KMgCl ₂ •6H ₂ O	0.00000	7819.3			
CaCl ₂ •2MgCl ₂ •12H ₂ O	0.00000	0.27626E+19			
Na ₂ SO ₄ •10H ₂ O	0.17043E-01	0.32600E-02			
Na ₂ SO ₄	0.00000	0.69007			
MgSO ₄ •6H ₂ O	0.00000	0.18782			
MgSO ₄ •7H ₂ O	0.00000	0.54675E-01			
K ₂ SO ₄	0.00000	0.69458E-02			
MgSO ₄ •K ₂ SO ₄ •6H ₂ O	0.00000	0.50713E-04			
CaSO ₄ (ion-pair)		0.60935E-01			
MgSO ₄ (ion-pair)		239.57			
Iterations = 3					

printed and the program exits. The most likely explanations for nonconvergence within 300 iterations are divergence, oscillating around the correct solution, or the system is beyond the eutectic, at which point only solid phases are thermodynamically stable.

To rectify mathematical problems, try changing the temperature or evaporative increments, either increasing or decreasing their step size (e.g., change temperature decrement from 10 to 5 degrees or change water decrement from 50 to 100 g).

Generally one can tell the approximate eutectic of more complex mixtures from the behavior of pure salts. For example, the eutectic temperature for pure CaCl₂ is -50.4 °C (Spencer et al. 1990). Failure of the model to converge at -40 °C for a

solution containing CaCl₂ such as seawater cannot be due to exceeding the eutectic, which must be lower in temperature than the eutectic of the pure solution. Failure of the model to converge within 300 iterations at -54 °C for seawater is caused by exceeding the eutectic at -53.6 °C (Spencer et al. 1990).

If you use FREZCHEM to identify "junctions" such as the eutectic, you need to carefully scrutinize the output for accuracy. For example, FREZCHEM prints results for seawater freezing down to -53.7 °C; at -53.8 °C, the program fails to converge within 300 iterations because the solution is beyond the eutectic temperature. If one examines the printed output at -53.7 °C, the solution is slightly supersaturated with respect to CaCl₂

(the last of the seawater salts to precipitate) based on calculated activity products, but the program indicates no precipitation of CaCl_2 (moles = 0.0). At -53.6°C , the solution is undersaturated with respect to CaCl_2 . The true eutectic lies between -53.7°C and -53.6°C , which agrees with Spencer et al. (1990) who place the seawater eutectic at -53.64°C . Small errors can occur at junctions; therefore, these solutions need to be carefully scrutinized.

FREZCHEM is limited to applications where chloride and sulfate salts of sodium, potassium, calcium, and magnesium are dominant. Two common salts not currently in the model are gypsum [$\text{CaSO}_4 \cdot 2\text{H}_2\text{O}(\text{cr})$] and calcite [$\text{CaCO}_3(\text{cr})$]. In the evaporative scenario (Table 4), the calculated activity product for gypsum in seawater is 2.09×10^{-4} which is apparently supersaturated with respect to gypsum at 0°C ($K = 2.33 \times 10^{-5}$, Plummer et al. 1988). The reason gypsum and calcite are missing from the current model is because these salts are relatively insoluble and as a consequence, there are few solubility data at subzero temperatures. Because neither gypsum nor calcite can be parameterized over the temperature range of interest (-60 to $+25^\circ\text{C}$), we have chosen not to include their solubility, even over a limited temperature range,

in this preliminary version of FREZCHEM.

This is not to say that FREZCHEM cannot be used in circumstances where gypsum or calcite are precipitating, only that caution is necessary in interpreting the results. For example, calcite precipitates during both seawater freezing (Nelson and Thompson 1954) and seawater evaporation, as does gypsum during seawater evaporation (McCaffrey et al. 1987). However, the amounts of calcite precipitation are relatively minor and can apparently be ignored. Note the good agreement in Ca concentrations during seawater freezing between experimental data and the model that ignored both calcite and gypsum (Figs. 3 and 4). One cannot ignore gypsum during seawater evaporation (McCaffrey et al. 1987, Herut et al. 1990) where explicit recognition must be made of gypsum precipitation. As pointed out previously, a 10-fold concentration of seawater at 0°C is likely to lead to the significant precipitation of both mirabilite (0.0170 moles) and gypsum (supersaturated) (Table 4). Where gypsum precipitation is possible, check the calculated gypsum activity product for apparent supersaturation. Unfortunately, reliable gypsum solubility products are available only for temperatures $\geq 0^\circ\text{C}$ (Plummer et al. 1988).

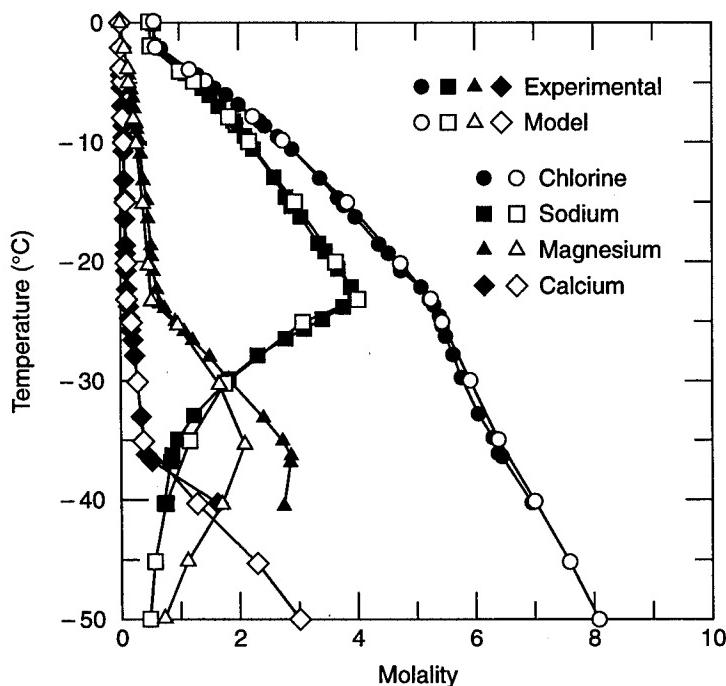


Figure 3. The concentrations of major seawater constituents during freezing. Experimental measurements are from Nelson and Thompson (1954).

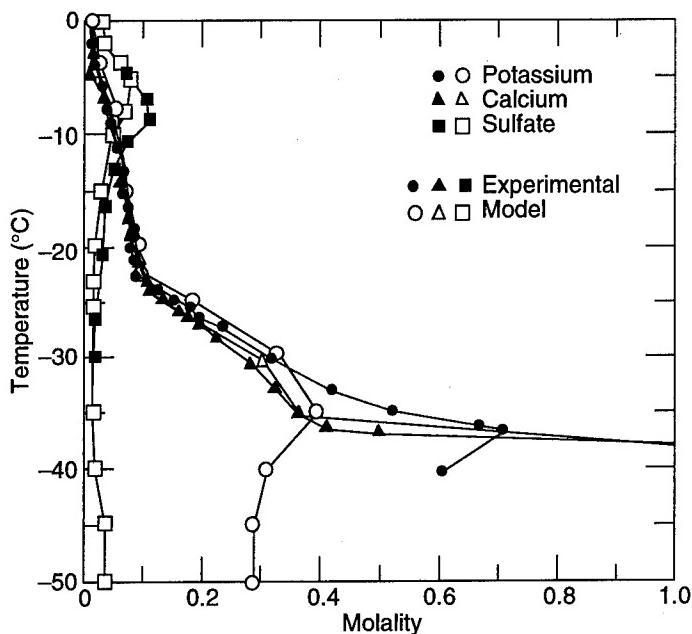


Figure 4. The concentrations of minor seawater constituents during freezing. Experimental measurements are from Nelson and Thompson (1954).

PROGRAM VALIDATION

The Spencer-Møller-Weare model and consequently the FREZCHEM model are parameterized primarily with data from pure binary and ternary salt solutions. Spencer et al. (1990) present 13 figures demonstrating how well their model fits the experimental data (e.g., Fig. 1). In general, the agreement is excellent, which is not surprising because this is the database used to parameterize the model, and as a consequence, this is not a true validation of the model.

The primary validation of the Spencer-Møller-Weare model is a comparison of model predictions for seawater freezing with experimental data from Nelson and Thompson (1954). The comparison of major (Fig. 3) and minor (Fig. 4) seawater constituents demonstrates good agreement except for Mg and K concentrations between -30° and -40 °C and SO₄ concentrations between -5° and -10 °C. The seawater experimental data between -30° and -40 °C show a great excess of cations over anions indicating a major problem with the analytical measurements. All of the discrepancy between measured and predicted Mg and K concentrations can be attributed to this charge imbalance. The model predicts that mirabilite precipitates at -5.9 °C, while experimental measurements suggest that precipitation begins at -8.2 °C. This is the

largest discrepancy in the temperature of first appearance of salt between the model and the experimental data. We plan to address these problems with experimental work in the near future.

Another approach to validating the model, not used by Spencer et al. (1990), is to examine the distribution of water during seawater freezing. For a system starting with 1.0 kg of water at 0 °C, approximately 50% of the water has turned to ice by the time the temperature has dropped to -4 °C (Fig. 5). By -10°C, only 20% of the original water is unfrozen. At about -23 °C, hydrated salts become a significant sink for water; this is where hydrohalite begins to precipitate. At -50 °C, the original 1000 g of water is redistributed as 964.73 g of ice, 32.12 g of hydrated salts, and 3.15 g of unfrozen water. There is good agreement between calculated unfrozen water and experimental measurements (Richardson 1976) at least down to -36 °C (Fig. 6). Below -36 °C, the experimental and model estimates diverge but come back together at -50 °C where the model predicts 3.15 g water and the experimental measurement indicates 3.53 g water. The two curves diverge at the point where MgCl₂•12H₂O(cr) begins to precipitate at -36 °C. Given major uncertainties in both model parameterization and experimental measurements, we cannot determine if the model or experimental measurements are more accurate at low temperature.

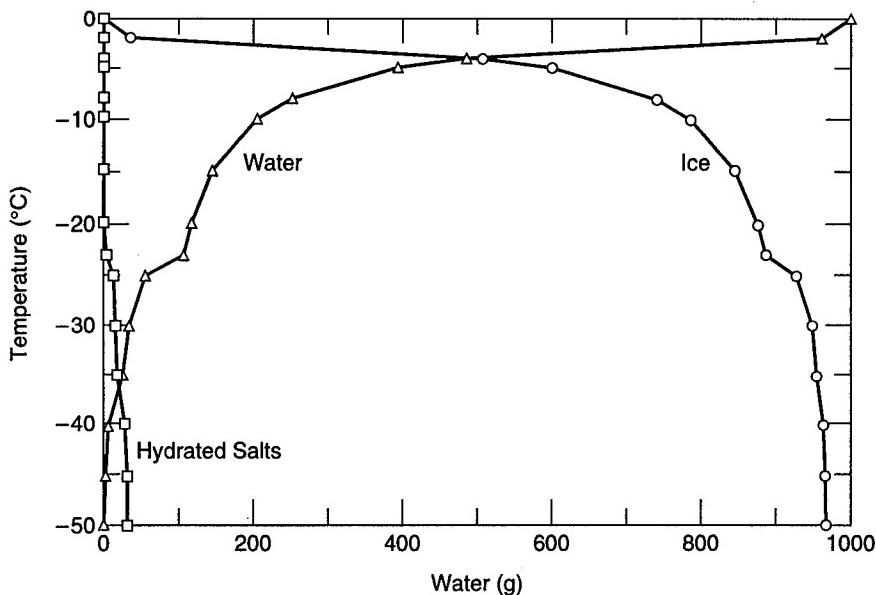


Figure 5. The model prediction of the distribution of water during seawater freezing.

The importance of salt precipitation in controlling seawater freezing is illustrated by examining the distribution of total Na and Cl during the freezing process (Fig. 7). At -50°C , only 0.3 % and 4.5 % of the original soluble Na and Cl at 0°C remain in the solution phase; 99.7 % and 95.5 % of the Na and Cl have precipitated in four salts.

The ionic strength of the unfrozen seawater changes from 0.72 to 11.85 molal between 0°C and -50°C . This large increase in salt concentration during the freezing process is the primary reason why it is necessary to use the Pitzer equations to evaluate the activity of water and activity coefficients.

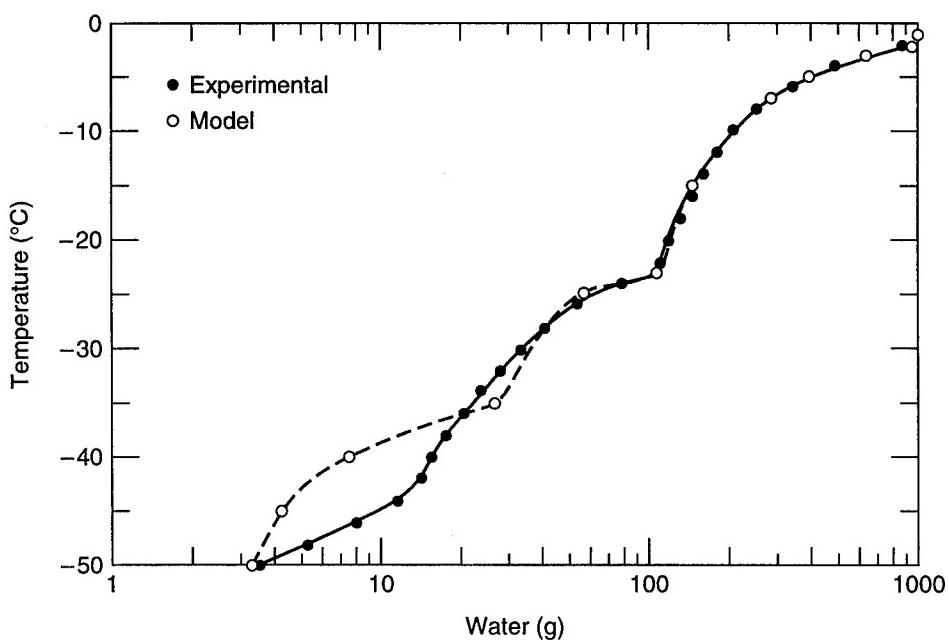


Figure 6. The amount of unfrozen water during seawater freezing. Experimental measurements are from Richardson (1976).

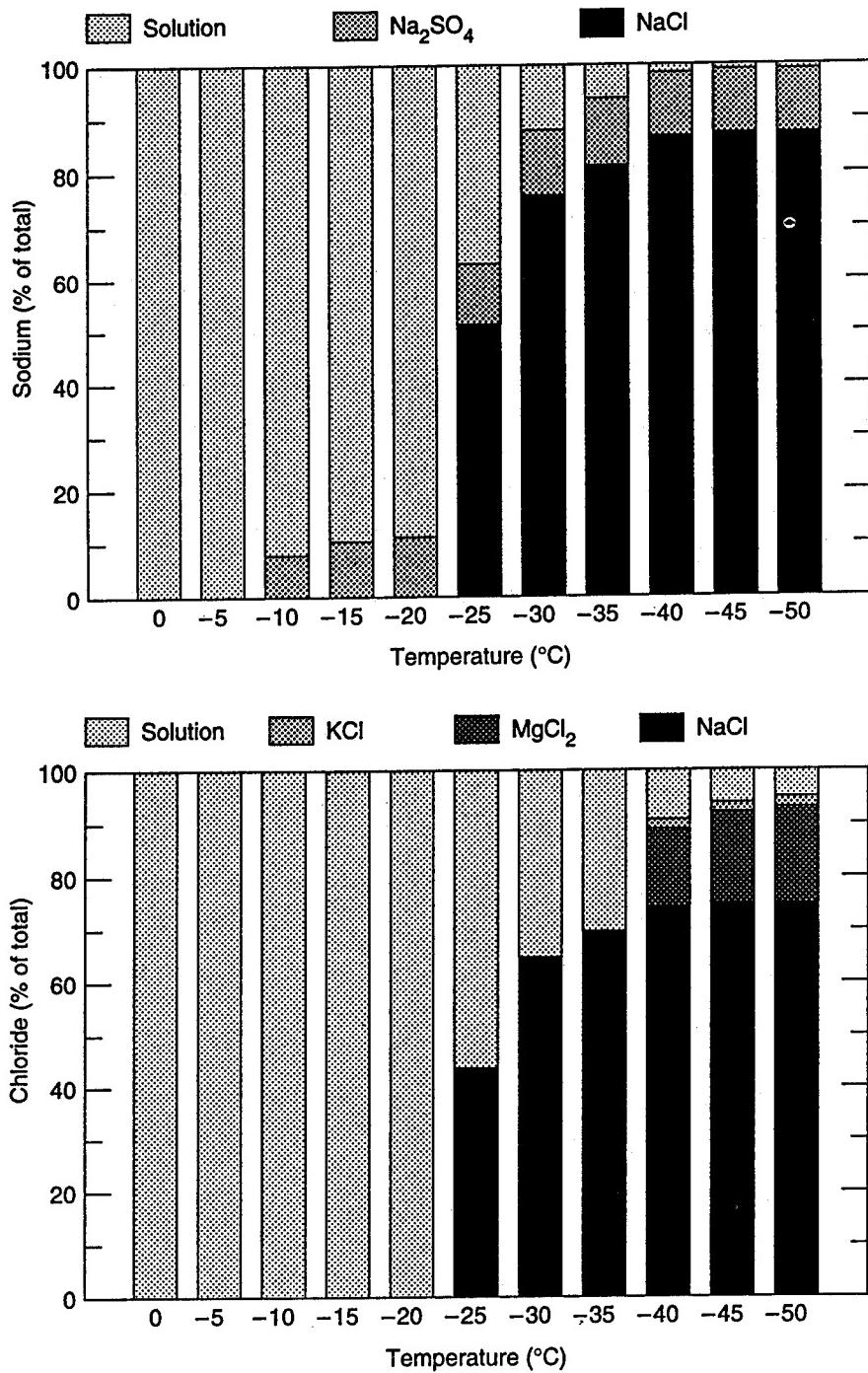


Figure 7. The model predictions of sodium and chloride distribution during seawater freezing.

FUTURE DIRECTIONS

The data presented in this report (Fig. 3-7) demonstrate the great complexity of the solute-freezing interaction and why this interaction must be considered to define accurately the freezing process, especially at high salinities. The

FREZCHEM model can calculate the role of the Cl and SO_4 salts of Na, K, Ca, and Mg in the freezing process. The model is useful as is, but more work needs to be done in order to develop a new "state-of-the-art" model.

For the chloride salts, KCl solubility is the major uncertainty. The parameterization of KCl solubil-

ity by Spencer et al. (1990) is based on only one data point below -24°C. We are beginning experiments to measure the solubility of KCl in KCl-CaCl₂ solutions to -50°C. At a minimum, we want to add gypsum [CaSO₄•2H₂O(cr)] solubility to the model and re-examine the solubility of Na₂SO₄, which is not predicted well by the current model (Fig. 4). And finally, we need to add calcite [CaCO₃(cr)] solubility to the model. Calcite and gypsum are common minerals that are present in many natural systems. Their inclusion in the model would be critical to many applications.

There are mathematical convergence problems, at times, with the FREZCHEM model, probably due to the sequential approach to solving a set of nonlinear equations used in the model. An alternative mathematical approach is the Gibbs energy minimization algorithm (Greenberg 1986, Harvie et al. 1987, Spencer et al. 1990). Our present strategy is to first document, by using the FREZCHEM model, where convergence problems exists. If these problems appear to be intractable using our present mathematical algorithm, we plan to try Gibbs energy minimization.

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Appendix A: FORTRAN listing of the FREZCHEM program.

4/6/94 15:25

Macintosh HD:MPW:FREZCHEM2.f

Page 1

PROGRAM FREZCHEM

C FREZCHEM: This is a chemical thermodynamic model for aqueous
C solutions over the temperature range from +25C to -60C. This
C model uses the Pitzer equations for activity coefficients and
C the activity of water and is valid to high ionic strengths
C (20 M). The model considers the precipitation-dissolution of
C the chloride and sulfate salts of calcium, magnesium, sodium,
C and potassium.

```
REAL IX,INIT(20),I,K
DIMENSION FINAL(20)
CHARACTER*50 TITLE
DOUBLE PRECISION AKI,AKII,ISALGM,SALGM,DELGM,SALH2O,X
DOUBLE PRECISION M,MO,H2O

COMMON /A/ MAXCAT,MAXAN,AKI(0:20),AKII(0:20),BET0(10,11:20),
xBET1(10,11:20),BET2(10,11:20),C(10,11:20),CPHI(10,11:20),
xPSI(20,20,20),THETA(20,20),A,Z(30)
COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
xIX(30),TITLE,ITER,AX(60),TEMP,I,PHI

OPEN (UNIT = 2, FILE = "FrData")
```

C Read in or initialize run parameters.

```
PRINT*, 'TITLE ='
READ(*,1) TITLE
1 FORMAT (A50)
PRINT*, 'FREEZE(1) OR EVAPORATION(2) SCENARIO ?'
READ*, PATH
PRINT*, 'SODIUM (M/KG) ='
READ*, IX(1)
PRINT*, 'POTASSIUM (M/KG) ='
READ*, IX(2)
PRINT*, 'CALCIUM (M/KG) ='
READ*, IX(3)
PRINT*, 'MAGNESIUM (M/KG) ='
READ*, IX(4)
PRINT*, 'CHLORIDE (M/KG) ='
READ*, IX(11)
PRINT*, 'SULFATE (M/KG) ='
READ*, IX(12)
PRINT*, 'INITIAL TEMPERATURE (K) ='
READ*, TEMP
IF (PATH.EQ.1) THEN
    PRINT*, 'FINAL TEMPERATURE (K) ='
    READ*, FINTEM
    PRINT*, 'TEMPERATURE DECREMENT (K) ='
    READ*, DELTEM
ELSE
    PRINT*, 'FINAL WATER (G) = '
    READ*, FINWAT
    PRINT*, 'WATER DECREMENT (G) = '
    READ*, DELWAT
END IF
DATA DEL / 1.25, 1.10, 1.01, 1.001, 1.0001 /
X(1) = IX(1)
X(2) = IX(2)
X(3) = IX(3)
X(4) = IX(4)
MAXCAT = 4
X(11) = IX(11)
```

```

X(12) = IX(12)
MAXAN = 12
X(30) = 1000
X(31) = 0
ISALGM = 0
3    ITER = 0

      CALL PARAMETER
5    CALL PITZER

C Test for supersaturation with respect to ice phase and freeze
C if necessary.

      IF (AH2O .GT. K(31)) THEN
          IF (PATH.EQ.2) THEN
              WRITE(2,*) 'THEORETICALLY THIS SOLUTION SHOULD FREEZE.'
              WRITE (2,*) 'THEREFORE, CAN NOT EVALUATE THIS USING AN
X EVAPORATIVE SCENARIO.'
              FINTEM=TEMP
              DELTEM=1.0
              PATH=1
          END IF
          MO = 0
          M = 0
          H2O = X(30) + X(31)
          DO 10 J = 1, 25
              MO = MO + X(J)
10        CONTINUE
          DO 70 KA = 1, 5
              DO 30 J = 1, 25
                  X(J) = X(J) * DEL(KA)
30        CONTINUE
              DO 40 J = 32, 60
                  X(J) = X(J) * DEL(KA)
40        CONTINUE
              CALL PITZER
              IF (AH2O .GT. K(31)) GOTO 20
              DO 50 J = 1, 25
                  X(J) = X(J) / DEL(KA)
50        CONTINUE
              DO 60 J = 32, 60
                  X(J) = X(J) / DEL(KA)
60        CONTINUE
              CONTINUE
              DO 80 J = 1, 25
                  M = M + X(J)
80        CONTINUE
              X(30) = X(30) * MO / M
              X(31) = H2O - X(30)
          END IF

C Test for mineral saturation and equilibrate solutions.

200    DO 210 J = 1, 20
            INIT(J) = X(J)
210    CONTINUE
            CALL PITZER
            ITER = ITER + 1
            IF ((IX(3) .GT. 0) .AND. (IX(12) .GT. 0)) CALL CASO4
            IF ((IX(4) .GT. 0) .AND. (IX(12) .GT. 0)) CALL MGSO4
            IF ((IX(1) .GT. 0) .AND. (IX(12) .GT. 0)) CALL NA2SO4
            IF ((IX(2) .GT. 0) .AND. (IX(12) .GT. 0)) CALL ARCANITE
            IF ((IX(2) .GT. 0) .AND. (IX(4) .GT. 0) .AND. (IX(12) .GT. 0))

```

```

x      CALL PICROMERITE
IF ((IX(2) .GT. 0) .AND. (IX(11) .GT. 0)) CALL SYLVITE
IF ((IX(1) .GT. 0) .AND. (IX(11) .GT. 0)) CALL NACL
IF ((IX(4) .GT. 0) .AND. (IX(11) .GT. 0)) CALL MGCL2
IF ((IX(3) .GT. 0) .AND. (IX(11) .GT. 0)) CALL ANTARCTICITE
IF ((IX(2) .GT. 0) .AND. (IX(4) .GT. 0) .AND. (IX(11) .GT. 0))
x      CALL CARNALLITE
IF ((IX(3) .GT. 0) .AND. (IX(4) .GT. 0) .AND. (IX(11) .GT. 0))
x      CALL TACHYHYDRITE
IF (ITER .EQ. 300) THEN
  WRITE (2,*) 'THIS RUN IS NOT CONVERGING AFTER 300
xINTERATIONS'
  GOTO 270
END IF

```

C Calculate water associated with hydrated salts.

```

SALH2O=2*X(32)+6*X(35)+6*X(36)+8*X(37)+12*X(38)+6*X(39)+12*X(40)
SALH2O=SALH2O+10*X(41)+6*X(43)+7*X(44)+6*X(46)
SALGM = SALH2O * X(30) * 18.0153 / 1000
DELGM = SALGM - ISALGM
X(30) = X(30) - DELGM
DO 220 J = 1, 25
  X(J) = X(J) * (X(30) + DELGM) / X(30)
220  CONTINUE
DO 230 J = 32, 60
  X(J) = X(J) * (X(30) + DELGM) / X(30)
230  CONTINUE
ISALGM = SALGM

```

C Test for mineral equilibria.

```

235  DO 240 J = 1, 20
      FINAL(J) = X(J)
      IF (FINAL(J) .EQ. 0) GOTO 240
      IF ((ABS(FINAL(J)-INIT(J))/FINAL(J))*1000 .GT. 1) GOTO 200
240  CONTINUE

```

C Test for ice-water equilibrium

```

CALL PITZER
IF (AH2O-K(31) .GT. 0.00005) GOTO 5

```

C Print Results.

260 CALL PPRINT

C Update Temperature or Water Content

```

IF (PATH.EQ.1) THEN
  TEMP = TEMP - DELTEM
  IF (TEMP .GE. FINTEM) GOTO 3
ELSE
  X(30)=X(30)-DELWAT
  TH2O = SALGM+X(30)
  IF (TH2O.LT.FINWAT) GOTO 270
  DO 263 IJK=1,25
    X(IJK)=X(IJK)*(X(30)+DELWAT)/X(30)
263  CONTINUE
  DO 265 IJK=32,60
    X(IJK)=X(IJK)*(X(30)+DELWAT)/X(30)
265  CONTINUE
  GO TO 3

```

```

    END IF
270   END
C-----
C-----

```

SUBROUTINE PARAMETER

C This subroutine calculates the parameters and their temperature
 C dependence for the Pitzer equations and mineral solubility products.

```

REAL IX,I,K
CHARACTER*50 TITLE
DOUBLE PRECISION D(64,6), P(64),AKI,AKII,T,X

COMMON /A/ MAXCAT,MAXAN,AKI(0:20),AKII(0:20),BETO(10,11:20),
xBET1(10,11:20),BET2(10,11:20),C(10,11:20),CPHI(10,11:20),
xPSI(20,20,20),THETA(20,20),A,Z(30)
COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
xIX(30),TITLE,ITER,AX(60),TEMP,I,PHI

DATA AKI / 1.925154014814667, -0.060076477753119,
x      -0.029779077456514, -0.007299499690937,
x      0.000388260636404, 0.000636874599598,
x      0.000036583601823, -0.000045036975204,
x      -0.000004537895710, 0.000002937706971,
x      0.000000396566462, -0.00000020209617,
x      -0.000000025267769, 0.000000013522610,
x      0.000000001229405, -0.000000000821969,
x      -0.000000000050847, 0.000000000046333,
x      0.00000000001943, -0.00000000002563,
x      -0.000000000010991 /

DATA AKII / 0.628023320520852, 0.462762985338493,
x      0.150044637187895, -0.028796057604906,
x      -0.036552745910311, -0.001668087945272,
x      0.006519840398744, 0.001130378079086,
x      -0.000887171310131, -0.000242107641309,
x      0.000087294451594, 0.000034682122751,
x      -0.000004583768938, -0.000003548684306,
x      -0.000000250453880, 0.000000216991779,
x      0.000000080779570, 0.000000004558555,
x      -0.000000006944757, -0.000000002849257,
x      0.000000000237816 /

DATA Z / 1,1,2,2,1,0,0,0,0,0,-1,-2,-1,-1,-2,
x      0,0,0,0,0,0,0,0,0,0,0,0,0,0,0,0,0,0,0,0,0 /

```

C Data Stored in D(I,J) are from Spencer et al. (1990)

```

DATA ((D(I,J),J=1,6),I=1,64) /
x      86.6836498, 0.0848795942, -8.8878515D-5,
x      4.88096393D-8, -1.32731477D3, -17.6460172,
x      7.87239712, -8.3864096D-3, 1.44137774D-5,
x      -8.7820301D-9, -496.920671, -0.820972560,
x      866.915291, 0.606166931, -4.8048921D-4,
x      1.88503857D-7, -1.70460145D4, -167.171296,
x      1.70761824, 2.32970177D-3, -2.46665619D-6,
x      1.21543380D-9, -1.35583596, -0.387767714,
x      26.5718766, 9.92715099D-3, -3.62323330D-6,
x      -6.28427180D-11, -755.707220, -4.67300770,

```

```

x      1.69742977D3, 1.22270943, -9.99044490D-4,
x      4.04786721D-7, -3.28684422D4, -328.813848,
x      -3.27571680, -1.27222054D-3, 4.71374283D-7,
x      1.1162507D-11, 90.7747666, 0.580513562,
x      -5.62764702D1, -3.00771997D-2, 1.05630400D-5,
x      3.3331626D-9, 1.11730349D3, 1.06664743D1,
x      3.4787, -1.5417D-2, 3.1791D-5, 0, 0, 0,
x      2.64231655D1, 2.46922993D-2, -2.48298510D-5,
x      1.22421864D-8, -4.18098427D2, -5.35350322,
x      3.13852913D2, 2.61769099D-1, -2.46268460D-4,
x      1.15764787D-7, -5.53133381D3, -6.21616862D1,
x      -3.18432525D4, -2.86710358D1, 2.78892838D-2,
x      -1.3279705D-5, 5.24032958D5, 6.40770396D3,
x      5.9532D-2, -2.49949D-4, 2.41831D-7, 0, 0, 0,
x      -3.32486330D3, -2.9297353, 2.8024367D-3,
x      -1.316883D-6, 5.53958527D4, 6.66660369D2,
x      -3.57406160D3, -3.0011206, 2.73660950D-3,
x      -1.219171D-6, 6.09716482D4, 7.11613120D2,
x      3.68520478D2, 3.16243995D-1, -2.95372760D-4,
x      1.35491104D-7, -6.22607913D3, -7.35844094D1,
x      4.07908797D1, 8.26906675D-3, 0, 0, -1.41842998D3, -6.74728848,
x      -1.31669651D1, 2.35793239D-2, 0, 0, 2.06712594D3, 0,
x      -1.88D-2, 0, 0, 0, 0, 0,
x      1.50D-1, 0, 0, 0, 0, 0,
x      3.00, 0, 0, 0, 0, 0,
x      -1.29399287D2, 4.00431027D-1, 0, 0, 0, 0,
x      5.40007849D3, 4.90576884, -4.80489750D-3,
x      2.31126994D-6, -8.80664146D4, -1.08839565D3,
x      2.78730869, 4.30077440D-3, 0, 0, 0, 0,
x      -5.88623653D2, -5.0552280D-1, 4.8277657D-4,
x      -2.3029838D-7, 1.02002016D4, 1.17303808D2,
x      -18.2266741, -3.69038470D-3, 0,
x      0, 612.415011, 3.02994981,
x      6.48108127, 1.46803468D-3, 0,
x      0, -204.354019, -1.09448043,
x      3.48120D-2, 0, 0, 0, -8.21660, 0,
x      5.0D-2, 0, 0, 0, 0,
x      -7.63980, -1.2990D-2, 1.1060D-5, 0, 0, 1.8475,
x      -1.20D-2, 0, 0, 0, 0, 0,
x      7.0D-2, 0, 0, 0, 0, 0,
x      -3.109870D-2, 5.4464780D-5, 0, 0, 1.99404210, 0,
x      1.1750520D-1, 0, 0, 0, -4.198620D1, 0,
x      2.365710, -4.540D-3, 0, 0, -2.84940D2, 0,
x      -5.930D-2, 2.54280D-4, 0, 0, -1.34390D1, 0,
x      1.1670D-1, 0, 0, 0, 0, 0,
x      5.0362230D-2, -8.750820D-6, 0, 0, -2.899090D1, 0,
x      -1.3679157, 4.24016653D-3, 0, 0, 0, 0,
x      5.31274136, -6.3424248D-3, 0, 0, -9.83113847D2, 0,
x      4.15790220D1, 1.30377312D-2, 0, 0, -9.81658526D2, -7.4061986,
x      7.00D-2, 0, 0, 0, 0, 0,
x      4.0209775, 1.1286005D-3, 0, 0, -1.01169260D2, -7.060798D-1,
x      -2.124815D-1, 2.8469833D-4, 0, 0, 3.75619614D1, 0,
x      -1.80D-2, 0, 0, 0, 0, 0,
x      -1.839158D-1, 1.429444D-4, 0, 0, 3.263D1, 0,
x      -1.62917341D3, -1.51940390, 1.45249679D-3,
x      -6.9427505D-7, 2.26012743D4, 333.075506,
x      7.875060393D3, 11.69118490, -1.7183789D-2,
x      1.24395543D-5, -9.3314790D4, -1.7287461D3,
x      -1.2222551D4, -9.8806459, 8.46685083D-3,
x      -3.4459117D-6, 2.09823965D5, 2.42328528D3,
x      -1.47477745D1, 0, 0, 0, 3.26409496D3, 0,
x      3.65283115D2, -7.1578257D-1, 0, 0, -4.48753391D4, 0,
x      9.14839001D3, 8.22348745, -8.1288759D-3,

```

```

x      3.95552403D-6, -1.54040868D5, -1.83624247D3,
x      1.42290062D5, 1.61973105D2, -1.95332071D-1,
x      1.17636119D-4, -2.04059847D6, -2.97464810D4,
x      7.52225099D2, 1.17584653D-1, 0, 0, -2.43223909D4, -1.21990076D2,
x      2.27801976D3, 6.49361616D-1, 0, 0, -6.23075123D4, -3.95438891D2,
x      2.55008896D5, 2.44532240D2, -2.48807876D-1,
x      1.22425236D-4, -4.02988342D6, -5.18668604D4,
x      -4.45702171D1, 2.32023790D-1, -7.14935692D-4,
x      5.32658215D-7, -4.24817923D3, 8.59110245,
x      8.03777918D1, -1.388069D-1, 0, 0, 0,
x      8.44728050D4, 7.68443387D1, -7.4825816D-2,
x      3.51806085D-5, -1.3881852D6, -1.7026778D4,
x      -3.9635632D3, -5.8114490, 7.59799462D-3,
x      -4.6571737D-6, 3.93454893D4, 8.79598423D2,
x      2.39166461D1, -9.3680744D-2, 0, 0, 0, 0,
x      2.02443627, -1.8051586D-2, 0, 0, 0, 0,
x      -3.53346251D4, -2.76809991D1, 2.39856021D-2,
x      -1.0201845D-5, 6.37561988D5, 6.94737787D3,
x      -1.75203509D1, 2.793713D-2, 0, 0, 0, 0 /
T=TEMP
DO 730 J = 1, 64
    P(J)=D(J,1)+D(J,2)*T+D(J,3)*T**2+D(J,4)*T**3+D(J,5)
    x          /T+D(J,6)*LOG(T)

```

730 CONTINUE

```

A = P(1)
BETO(1,11) = P(2)
BET1(1,11) = P(3)
CPHI(1,11) = P(4)
BETO(2,11) = P(5)
BET1(2,11) = P(6)
CPHI(2,11) = P(7)
BETO(3,11) = P(8)
BET1(3,11) = P(9)
CPHI(3,11) = P(10)
BETO(4,11) = P(11)
BET1(4,11) = P(12)
CPHI(4,11) = P(13)
BETO(1,12) = P(14)
BET1(1,12) = P(15)
CPHI(1,12) = P(16)
BETO(2,12) = P(17)
BET1(2,12) = P(18)
CPHI(2,12) = P(19)
BETO(3,12) = P(20)
BET1(3,12) = P(21)
BET2(3,12) = P(22)
BETO(4,12) = P(23)
BET1(4,12) = P(24)
CPHI(4,12) = P(25)
THETA(1,2) = P(26)
PSI(1,2,11) = P(27)
PSI(1,2,12) = P(28)
THETA(1,3) = P(29)
PSI(1,3,11) = P(30)
PSI(1,3,12) = P(31)
THETA(1,4) = P(32)
PSI(1,4,11) = P(33)
PSI(1,4,12) = P(34)
THETA(2,3) = P(35)
PSI(2,3,11) = P(36)
THETA(2,4) = P(37)
PSI(2,4,11) = P(38)

```

```

PSI(2,4,12) = P(39)
THETA(3,4) = P(40)
PSI(3,4,11) = P(41)
PSI(3,4,12) = 0.024
THETA(11,12) = P(42)
PSI(11,12,1) = P(43)
PSI(11,12,2) = P(44)
PSI(11,12,3) = P(45)
PSI(11,12,4) = P(46)
K(34) = EXP(P(47))
K(31) = EXP(P(48))
K(32) = EXP(P(49))
K(22) = EXP(P(50))
K(23) = EXP(P(51))
K(33) = EXP(P(52))
K(35) = EXP(P(53))
K(36) = EXP(P(54))
K(37) = EXP(P(55))
K(38) = EXP(P(56))
K(39) = EXP(P(57))
K(40) = EXP(P(58))
K(41) = EXP(P(59))
K(42) = EXP(P(60))
K(43) = EXP(P(61))
K(44) = EXP(P(62))
K(45) = EXP(P(63))
K(46) = EXP(P(64))
RETURN
END

```

C-----
C-----

SUBROUTINE PITZER

C This subroutine calculates activity coefficients and the activity
C of water using the Pitzer equations.

```

REAL I,IX,K
CHARACTER*50 TITLE
DOUBLE PRECISION AKI,AKII,X
DIMENSION SUMCA(10), SUMCAT(10), SUMAN(10), SUMZ(10),
x      SUMAC(11:20), SUMAA(11:20), SUMCC(11:20), SUMK(11:20),
x      BPHI(10,11:20), B(10,11:20), BPRIME(10,11:20),
x      PHIPHI(20,20), PHIPRI(20,20), PHIIJ(20,20)

COMMON /A/ MAXCAT,MAXAN,AKI(0:20),AKII(0:20),BETO(10,11:20),
xBET1(10,11:20),BET2(10,11:20),C(10,11:20),CPHI(10,11:20),
xPSI(20,20,20),THETA(20,20),A,Z(30)
COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
xIX(30),TITLE,ITER,AX(60),TEMP,I,PHI

```

```

ZZ = 0
SM = 0
I = 0
DO 280 J = 1, 20
    ZZ = ZZ + X(J) * ABS(Z(J))
    I = I + .5 * (X(J) * Z(J)**2)
    SM = SM + X(J)

```

280 CONTINUE

DATA SUMCA / 10*0 /

```

DATA SUMCAT / 10*0 /
DATA SUMAN / 10*0 /
DATA SUMZ / 10*0 /
DATA SUMAC / 10*0 /
DATA SUMAA / 10*0 /
DATA SUMCC / 10*0 /
DATA SUMK / 10*0 /

ALPHA = 2 * SQRT(I)
ALPHA1 = 1.4 * SQRT(I)
ALPHA2 = 12 * SQRT(I)
G1 = 2*(1-(1+ALPHA1)*EXP(-ALPHA1))/ALPHA1**2
G2 = 2*(1-(1+ALPHA2)*EXP(-ALPHA2))/ALPHA2**2
GPRI1 = -2*(1-(1+ALPHA1+ALPHA1**2/2)*EXP(-ALPHA1))/ALPHA1**2
GPRI2 = -2*(1-(1+ALPHA2+ALPHA2**2/2)*EXP(-ALPHA2))/ALPHA2**2
G = 2*(1-(1+ALPHA)*EXP(-ALPHA))/ALPHA**2
GPRIME = -2*(1-(1+ALPHA+ALPHA**2/2)*EXP(-ALPHA))/ALPHA**2
DO 300 J = 1, MAXCAT
    DO 290 IK = 11, MAXAN
        IF (Z(J)*ABS(Z(IK)) .EQ. 4) THEN
            BPHI(J,IK)=BETO(J,IK)+BET1(J,IK)*EXP(-ALPHA1) +
x           BET2(J,IK)*EXP(-ALPHA2)
            B(J,IK)=BETO(J,IK)+BET1(J,IK)*G1+BET2(J,IK)*G2
            BPRIME(J,IK)=BET1(J,IK)*GPRI1/I+BET2(J,IK)*
x           GPRI2/I
        ELSE
            BPHI(J,IK)=BETO(J,IK)+BET1(J,IK)*EXP(-ALPHA)
            B(J,IK)=BETO(J,IK)+BET1(J,IK)*G
            BPRIME(J,IK)=BET1(J,IK)*GPRIME/I
        END IF
        C(J,IK)=CPHI(J,IK)/(2*SQRT(Z(J)*ABS(Z(IK))))
290    CONTINUE
300    CONTINUE
    DO 320 J = 1, MAXCAT - 1
        DO 310 IK = J + 1, MAXCAT
            IF ((Z(J) .EQ. 0) .OR. (Z(IK) .EQ. 0)) GOTO 310
            CALL INTERACTION(Z(J),Z(IK),I,A,PHIPHI(J,IK),
x           PHIPRI(J,IK),PHIIJ(J,IK),THETA(J,IK),AKI,AKII)
310    CONTINUE
320    CONTINUE
    DO 340 J = 11, MAXAN - 1
        DO 330 IK = J + 1, MAXAN
            IF ((Z(J) .EQ. 0) .OR. (Z(IK) .EQ. 0)) GOTO 330
            CALL INTERACTION(Z(J),Z(IK),I,A,PHIPHI(J,IK),
x           PHIPRI(J,IK),PHIIJ(J,IK),THETA(J,IK),AKI,AKII)
330    CONTINUE
340    CONTINUE

```

C Calculation of summation terms for F and PHI.

```

SCATON = 0
SUBSUM = 0
SANON = 0
SUMCAF = 0
SUMANF = 0

DO 370 J = 1, MAXCAT - 1
    DO 360 IK = J + 1, MAXCAT
        DO 350 L = 11, MAXAN
            SUBSUM = SUBSUM + PSI(J,IK,L) * X(L)
350    CONTINUE
            SCATON=SCATON+(SUBSUM+PHIPHI(J,IK))*X(J)*X(IK)
            SUMCAF=SUMCAF+PHIPRI(J,IK)*X(J)*X(IK)

```

```

        SUBSUM = 0
360      CONTINUE
370      CONTINUE
        SUBSUM = 0
        DO 400 J = 11, MAXAN - 1
            DO 390 IK = J + 1, MAXAN
                DO 380 L = 1, MAXCAT
                    SUBSUM=SUBSUM+PSI (J,IK,L)*X(L)
380      CONTINUE
        SANON=SANON+ (SUBSUM+PHIPHI (J,IK))*X(J)*X(IK)
        SUMANF=SUMANF+PHIPRI (J,IK)*X(J)*X(IK)
        SUBSUM = 0
390      CONTINUE
400      CONTINUE
        SUMB = 0
        SUMPHI = 0
        DO 420 J = 1, MAXCAT
            DO 410 IK = 11, MAXAN
                SUMB=SUMB+X(J)*X(IK)*BPRIME (J,IK)
                SUMPHI=SUMPHI+X(J)*X(IK)*(BPHI (J,IK)+ZZ*C(J,IK))
410      CONTINUE
420      CONTINUE
        F==A*(SQRT(I)/(1+1.2*SQRT(I))+2*LOG(1+1.2*SQRT(I))/1.2)+SUMB+
        X     SUMCAF+SUMANF
        PHI=1+(2/SM)*((-A*I**1.5)/(1+1.2*SQRT(I))+SUMPHI+SCATON+
        X     SANON)
        AH2O = EXP(-PHI*SM/55.50837)

```

C Calculation of terms for activity coefficients (gamma).

```

        SUM = 0
        DO 450 J = 1, MAXCAT - 1
            DO 440 IK = J + 1, MAXCAT
                DO 430 L = 11, MAXAN
                    PSI(IK,J,L) = PSI(J,IK,L)
                    PHIIJ(IK,J) = PHIIJ(J,IK)
430      CONTINUE
440      CONTINUE
450      CONTINUE
        DO 480 L = 11, MAXAN - 1
            DO 470 IK = L + 1, MAXAN
                DO 460 J = 1, MAXCAT
                    PSI(IK,L,J) = PSI(L,IK,J)
                    PHIIJ(IK,L) = PHIIJ(L,IK)
460      CONTINUE
470      CONTINUE
480      CONTINUE
        DO 500 J = 1, MAXCAT
            DO 490 IK = 11, MAXAN
                SUMCA(J) = SUMCA(J)+X(IK)*(2*(B(J,IK))+ZZ*C(J,IK))
490      CONTINUE
500      CONTINUE
        DO 530 J = 1, MAXCAT
            DO 520 IK = 1, MAXCAT
                IF (J .EQ. IK) GOTO 520
                DO 510 L = 11, MAXAN
                    SUM=SUM+X(L)*PSI (J,IK,L)
510      CONTINUE
                SUMCAT(J)=SUMCAT(J)+X(IK)*(SUM+2*PHIIJ(J,IK))
                SUM = 0
520      CONTINUE
530      CONTINUE
        DO 560 J = 1, MAXCAT

```

```

      DO 550 IK = 11, MAXAN-1
          DO 540 L = IK + 1, MAXAN
              SUMAN(J) = SUMAN(J)+X(IK)*X(L)*PSI(IK,L,J)
540      CONTINUE
550      CONTINUE
560      CONTINUE
      SUM = 0
      DO 580 J = 1, MAXCAT
          DO 570 IK = 11, MAXAN
              SUM = SUM + X(J)*X(IK)*C(J,IK)
570      CONTINUE
580      CONTINUE
      DO 590 J = 1, MAXCAT
          SUMZ(J) = SUM*ABS(Z(J))
590      CONTINUE
      DO 600 J = 1, MAXCAT
          GX(J)=EXP(Z(J)**2*F+SUMCA(J)+SUMCAT(J)+SUMAN(J)+SUMZ(J))
600      CONTINUE
      SUM = 0
      DO 620 IK = 11, MAXAN
          DO 610 J = 1, MAXCAT
              SUMAC(IK) = SUMAC(IK)+X(J)*(2*B(J,IK)+ZZ*C(J,IK))
610      CONTINUE
620      CONTINUE
      DO 650 IK = 11, MAXAN
          DO 640 L = 11, MAXAN
              IF (IK .EQ. L) GOTO 640
              DO 630 J = 1, MAXCAT
                  SUM = SUM+X(J)*PSI(IK,L,J)
630      CONTINUE
      SUMAA(IK)=SUMAA(IK)+X(L)*(2*PHIIJ(IK,L)+SUM)
      SUM = 0
640      CONTINUE
650      CONTINUE
      DO 680 IK = 11, MAXAN
          DO 670 J = 1, MAXCAT - 1
              DO 660 L = J + 1, MAXCAT
                  SUMCC(IK)=SUMCC(IK)+X(J)*X(L)*PSI(J,L,IK)
660      CONTINUE
670      CONTINUE
680      CONTINUE
      SUM = 0
      DO 700 J = 1, MAXCAT
          DO 690 IK = 11, MAXAN
              SUM = SUM+X(J)*X(IK)*C(J,IK)
690      CONTINUE
700      CONTINUE
      DO 710 IK = 11, MAXAN
          SUMK(IK) = SUM * ABS(Z(IK))
710      CONTINUE
      DO 720 IK = 11, MAXAN
          GX(IK)=EXP(Z(IK)**2*F+SUMAC(IK)+SUMAA(IK)+SUMCC(IK)+SUMK(IK))
720      CONTINUE
      RETURN
      END

```

C-----

C-----

SUBROUTINE INTERACTION(Z1,Z2,I,A,PHIPHI,PHIPRI,PHIIJ,THETA,
xAKI,AKII)

C This subroutine calculates the higher-order electrostatic
C interaction terms for the Pitzer equations.

```

REAL I, JX(2,2), JXPRIME(2,2)
DIMENSION B(0:22), D(0:22), ZA(2), XA(2,2)
DOUBLE PRECISION AKI(0:20), AKII(0:20)

B(21) = 0
B(22) = 0
D(21) = 0
D(22) = 0
ZA(1) = Z1
ZA(2) = Z2

DO 770 JJ = 1, 2
    DO 760 IJ = 1, 2
        XA(JJ,IJ)=6*ZA(IJ)*ZA(JJ)*A*SQRT(I)
        X=XA(JJ,IJ)
        IF (X .LT. 1) THEN
            ZZ = .4*X**0.20-2.0
            DZ = 0.80*X**(-0.80)
            DO 740 K = 20, 0, -1
                B(K) = ZZ*B(K+1)-B(K+2)+AKI(K)
                D(K) = B(K+1)+ZZ*D(K+1)-D(K+2)
740      CONTINUE
        ELSE
            ZZ = (40/9)*X**(-0.10)-22/9
            DZ = (-40/90)*X**(-1.10)
            DO 750 K = 20, 0, -1
                B(K) = ZZ*B(K+1)-B(K+2)+AKII(K)
                D(K) = B(K+1)+ZZ*D(K+1)-D(K+2)
750      CONTINUE
        END IF
        JX(JJ,IJ) = .25*X-1+.5*(B(0)-B(2))
        JXPRIME(JJ,IJ) = .25+.5*DZ*(D(0)-D(2))
760      CONTINUE
770      CONTINUE
ETHETA=(Z1*Z2/(4*I))*(JX(1,2)-.5*JX(1,1)-.5*JX(2,2))
ETHPRI=-ETHETA/I+(Z1*Z2/(8*I**2))*(XA(1,2)*JXPRIME(1,2)-.5*
X     XA(1,1)*JXPRIME(1,1)-.5*XA(2,2)*JXPRIME(2,2))
PHIPHI=THETA+ETHETA+I*ETHPRI
PHIIJ=THETA+ETHETA
PHIPRI=ETHPRI
RETURN
END

```

C-----

C-----

SUBROUTINE PPRINT

C Subroutine prints the results.

```

REAL MOLE, K, IX, I
CHARACTER*20 SPECIS(30), SOLID(31:60)
CHARACTER*50 TITLE
DIMENSION AX(60)
DOUBLE PRECISION X

COMMON /B/ X(60), GX(60), K(20:60), AH2O, DEL(5),
XIX(30), TITLE, ITER, AX(60), TEMP, I, PHI

```

C

c

SUBROUTINE NACL

C Subroutine for NaCl equilibrium.

```

REAL IX,I,K
CHARACTER*50 TITLE
DOUBLE PRECISION NAT,CLT,KPRIME,X,DELTA

COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
XIX(30),TITLE,ITER,AX(60),TEMP,I,PHI

DELTA=0
HHALIT = K(32) / AH2O**2
HALITE = K(33)
NAT = X(1) + X(32) + X(33)
CLT = X(11) + X(32) + X(33)
IF (HHALIT .GT. HALITE) KPRIME = K(33)/(GX(1)*GX(11))
IF (HALITE .GT. HHALIT) KPRIME = K(32)/(GX(1)*GX(11)*AH2O**2)
IF (NAT*CLT-KPRIME .GT. 0) GOTO 860
855 X(1) = NAT
X(11) = CLT
X(32) = 0
X(33) = 0
GOTO 880
860 DECREM = MIN(NAT,CLT)
DO 870 KA = 2, 5
865   DELTA = DELTA+DECREM*(DEL(KA)-1)
      IF (DELTA .LT. 0) GOTO 855
      X(1) = NAT - DELTA
      X(11) = CLT - DELTA
      IF (KA .EQ. 5) CALL PITZER
      HHALIT = K(32) / AH2O**2
      HALITE = K(33)
      IF (HHALIT .GT. HALITE) KPRIME = K(33)/(GX(1)*GX(11))
      IF (HALITE .GT. HHALIT) KPRIME = K(32)/(GX(1)*GX(11)*
      X           AH2O**2)
      IF (X(1)*X(11)-KPRIME .GT. 0) GOTO 865
      DELTA = DELTA-DECREM*(DEL(KA)-1)
870 CONTINUE
      X(1) = NAT - DELTA
      X(11) = CLT - DELTA
      IF (HHALIT .GT. HALITE) THEN
          X(33) = DELTA
          X(32) = 0
      ELSE
          X(32) = DELTA
          X(33) = 0
      END IF
880 RETURN
END

```

C-----

C-----

SUBROUTINE SYLVITE

C Subroutine for KCl equilibrium.

```

REAL IX,I,K
CHARACTER*50 TITLE
DOUBLE PRECISION KT,CLT,KPRIME,X,DELTA

```

```

COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
XIX(30),TITLE,ITER,AX(60),TEMP,I,PHI

DELTA=0
KT = X(2) + X(34)
CLT = X(11) + X(34)
KPRIME = K(34) / (GX(2)*GX(11))
IF (KT*CLT-KPRIME .GT. 0) GOTO 890
885 X(2) = KT
X(11) = CLT
X(34) = 0
GOTO 910
890 DECREM = MIN(KT,CLT)
DO 900 KA = 2, 5
895     DELTA = DELTA+DECREM*(DEL(KA)-1)
     IF (DELTA .LT. 0) GOTO 885
     X(2) = KT - DELTA
     X(11) = CLT - DELTA
     IF (KA .EQ. 5) CALL PITZER
     KPRIME = K(34)/(GX(2)*GX(11))
     IF (X(2)*X(11)-KPRIME .GT. 0) GOTO 895
     DELTA = DELTA-DECREM*(DEL(KA)-1)
900 CONTINUE
X(2) = KT-DELTA
X(11) = CLT-DELTA
X(34) = DELTA
910 RETURN
END

```

C-----

C-----

SUBROUTINE ANTARCTICITE

C Subroutine for CaCl₂ equilibrium.

```

REAL IX,I,K
CHARACTER*50 TITLE
DOUBLE PRECISION CAT,CLT,KPRIME,X,DELTA

COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
XIX(30),TITLE,ITER,AX(60),TEMP,I,PHI

DELTA=0
CAT = X(3) + X(35)
CLT = X(11) + 2*X(35)
KPRIME = K(35)/(GX(3)*GX(11)**2*AH2O**6)
IF (CAT*CLT**2-KPRIME .GT. 0) GOTO 920
915 X(3) = CAT
X(11) = CLT
X(35) = 0
GOTO 940
920 DECREM = MIN(CAT,CLT)
DO 930 KA = 2, 5
925     DELTA = DELTA+DECREM*(DEL(KA)-1)
     IF (DELTA .LT. 0) GOTO 915
     X(3) = CAT - DELTA
     X(11) = CLT - 2*DELTA
     IF (KA .EQ. 5) CALL PITZER
     KPRIME = K(35)/(GX(3)*GX(11)**2*AH2O**6)
     IF (X(3)*X(11)**2-KPRIME .GT. 0) GOTO 925
     DELTA = DELTA - DECREM*(DEL(KA)-1)

```

```

930    CONTINUE
      X(3) = CAT-DELTA
      X(11) = CLT - 2*DELTA
      X(35) = DELTA
940    RETURN
      END

C-----
C-----
```

SUBROUTINE MGCL2

C Subroutine for MgCl₂ equilibrium.

```

REAL IX,I,K
CHARACTER*50 TITLE
DOUBLE PRECISION MGT,CLT,KPRIME,X,DELTA,KPRIM1,KPRIM2,KPRIM3

COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
           IX(30),TITLE,ITER,AX(60),TEMP,I,PHI

DELTA=0
MGT = X(4)+X(36)+X(37)+X(38)
CLT = X(11)+2*X(36)+2*X(37)+2*X(38)
KPRIM1 = K(36)/(GX(4)*GX(11)**2*AH2O**6)
KPRIM2 = K(37)/(GX(4)*GX(11)**2*AH2O**8)
KPRIM3 = K(38)/(GX(4)*GX(11)**2*AH2O**12)
KPRIME=MIN(KPRIM1,KPRIM2,KPRIM3)
IF (MGT*CLT**2-KPRIME .GT. 0) GOTO 960
950   X(4) = MGT
      X(11) = CLT
      X(36) = 0
      X(37) = 0
      X(38) = 0
      GOTO 990
960   DECREM = MIN(MGT,CLT)
      DO 980 KA = 2, 5
970       DELTA = DELTA+DECREM*(DEL(KA)-1)
      IF (DELTA .LT. 0) GOTO 950
      X(4) = MGT - DELTA
      X(11) = CLT-2*DELTA
      IF (KA .EQ. 5) CALL PITZER
      KPRIM1 = K(36)/(GX(4)*GX(11)**2*AH2O**6)
      KPRIM2 = K(37)/(GX(4)*GX(11)**2*AH2O**8)
      KPRIM3 = K(38)/(GX(4)*GX(11)**2*AH2O**12)
      KPRIME=MIN(KPRIM1,KPRIM2,KPRIM3)
      IF (X(4)*X(11)**2-KPRIME .GT. 0) GOTO 970
      DELTA = DELTA - DECREM*(DEL(KA)-1)
980   CONTINUE
      X(4) = MGT - DELTA
      X(11) = CLT - 2*DELTA
      IF (KPRIM1 .EQ. KPRIME) THEN
          X(36) = DELTA
          X(37) = 0
          X(38) = 0
      END IF
      IF (KPRIM2 .EQ. KPRIME) THEN
          X(37) = DELTA
          X(36) = 0
          X(38) = 0
      END IF
      IF (KPRIM3 .EQ. KPRIME) THEN

```

```

        X(38) = DELTA
        X(36) = 0
        X(37) = 0
990    END IF
        RETURN
        END

```

C-----

C-----

SUBROUTINE CARNALLITE

C Subroutine for KMgCl₃.6H₂O equilibrium.

```

REAL IX,I,K
CHARACTER*50 TITLE
DOUBLE PRECISION KT,MGT,CLT,KPRIME,X,DELTA

COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
XIX(30),TITLE,ITER,AX(60),TEMP,I,PHI

DELTA=0
KT = X(2) + X(39)
MGT = X(4) + X(39)
CLT = X(11) + 3*X(39)
KPRIME = K(39)/(GX(2)*GX(4)*GX(11)**3*AH2O**6)
IF (KT*MGT*CLT**3-KPRIME .GT. 0) GOTO 1005
1000  X(2) = KT
      X(11) = CLT
      X(4) = MGT
      X(39) = 0
      GOTO 1030
1005  DECREM = MIN(KT,MGT,CLT)
      DO 1020 KA = 2, 5
1010    DELTA = DELTA+DECREM*(DEL(KA)-1)
          IF (DELTA .LT. 0) GOTO 1000
          X(2) = KT - DELTA
          X(4) = MGT - DELTA
          X(11) = CLT - 3*DELTA
          IF (KA .EQ. 5) CALL PITZER
          KPRIME = K(39)/(GX(2)*GX(4)*GX(11)**3*AH2O**6)
          IF (X(2)*X(4)*X(11)**3-KPRIME .GT. 0) GOTO 1010
          DELTA = DELTA - DECREM*(DEL(KA)-1)
1020  CONTINUE
      X(2) = KT - DELTA
      X(4) = MGT - DELTA
      X(11) = CLT - 3*DELTA
      X(39) = DELTA
1030  RETURN
      END

```

C-----

C-----

SUBROUTINE TACHYHYDRITE

C Subroutine for CaMg₂Cl₆.12H₂O equilibrium.

```

REAL IX,I,K
CHARACTER*50 TITLE

```

```

COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
XIX(30),TITLE,ITER,AX(60),TEMP,I,PHI
      DOUBLE PRECISION KPRIME,CAT,MGT,CLT,X,DELTA

      DELTA=0
      CAT = X(3) + X(40)
      MGT = X(4) + 2*X(40)
      CLT = X(11) + 6*X(40)
      KPRIME = K(40)/(GX(3)*GX(4)**2*GX(11)**6*AH2O**12)
      IF (CAT*MGT**2*CLT**6-KPRIME .LT. 0) GOTO 1050
1040  X(3) = CAT
      X(11) = CLT
      X(4) = MGT
      X(40) = 0
      GOTO 1080
1050  DECREM = MIN(CAT,MGT,CLT)
      DO 1070 KA = 2, 5
1060  DELTA = DELTA+DECREM*(DEL(KA)-1)
      IF (DELTA .LT. 0) GOTO 1040
      X(3) = CAT - DELTA
      X(4) = MGT - 2*DELTA
      X(11) = CLT - 6*DELTA
      IF (KA .EQ. 5) CALL PITZER
      KPRIME = K(40)/(GX(3)*GX(4)**2*GX(11)**6*AH2O**12)
      IF (X(3)*X(4)**2*X(11)**6-KPRIME .GT. 0) GOTO 1060
      DELTA = DELTA-DECREM*(DEL(KA)-1)
1070  CONTINUE
      X(3) = CAT - DELTA
      X(4) = MGT - 2*DELTA
      X(11) = CLT - 6*DELTA
      X(40) = DELTA
1080  RETURN
      END

```

C-----

C-----

SUBROUTINE ARCANITE

C Subroutine for K2SO4 equilibrium.

```

REAL IX,I,K
CHARACTER*50 TITLE
DOUBLE PRECISION KT,SO4T,KPRIME,X,DELTA

COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
XIX(30),TITLE,ITER,AX(60),TEMP,I,PHI

      DELTA=0
      KT = X(2) + 2*X(45)
      SO4T = X(12) + X(45)
      KPRIME = K(45)/(GX(2)**2*GX(12))
      IF (KT**2*SO4T-KPRIME .GT. 0) GOTO 1100
1090  X(2) = KT
      X(12) = SO4T
      X(45) = 0
      GOTO 1130
1100  DECREM = MIN(KT,SO4T)
      DO 1120 KA = 2, 5
1110  DELTA = DELTA+DECREM*(DEL(KA)-1)
      IF (DELTA .LT. 0) GOTO 1090
      X(2) = KT - 2*DELTA

```

```

        X(12) = SO4T - DELTA
        IF (KA .EQ. 5) CALL PITZER
        KPRIME = K(45)/(GX(2)**2*GX(12))
        IF (X(2)**2*X(12)-KPRIME .GT. 0) GOTO 1110
        DELTA = DELTA - DECREM*(DEL(KA)-1)
1120 CONTINUE
        X(2) = KT - 2*DELTA
        X(12) = SO4T - DELTA
        X(45) = DELTA
1130 RETURN
END

```

C-----

C-----

SUBROUTINE NA2SO4

C Subroutine for Na2SO4 equilibrium.

```

REAL MIRAB,IX,I,K
CHARACTER*50 TITLE
DOUBLE PRECISION NAT,SO4T,KPRIME,X,DELTA

COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
           IX(30),TITLE,ITER,AX(60),TEMP,I,PHI

DELTA=0
NAT = X(1)+2*X(41)+2*X(42)
SO4T = X(12)+X(41)+X(42)
MIRAB = K(41)/AH2O**10
TENAR = K(42)
IF (MIRAB .GT. TENAR) THEN
    KPRIME = K(42)/(GX(1)**2*GX(12))
ELSE
    KPRIME = K(41)/(GX(1)**2*GX(12)*AH2O**10)
END IF
IF (NAT**2*SO4T-KPRIME .GT. 0) GOTO 1150
1140 X(1) = NAT
        X(12) = SO4T
        X(41) = 0
        X(42) = 0
        GOTO 1180
1150 DECREM = MIN(NAT,SO4T)
        DO 1170 KA = 2, 5
1160     DELTA = DELTA+DECREM*(DEL(KA)-1)
        IF (DELTA .LT. 0) GOTO 1140
        X(1) = NAT - 2*DELTA
        X(12) = SO4T - DELTA
        IF (KA .EQ. 5) CALL PITZER
        MIRAB = K(41)/AH2O**10
        TENAR = K(42)
        IF (MIRAB .GT. TENAR) THEN
            KPRIME = K(42)/(GX(1)**2*GX(12))
        ELSE
            KPRIME = K(41)/(GX(1)**2*GX(12)*AH2O**10)
        END IF
        IF (X(1)**2*X(12)-KPRIME .GT. 0) GOTO 1160
        DELTA = DELTA - DECREM*(DEL(KA)-1)
1170 CONTINUE
        X(1) = NAT - 2*DELTA
        X(12) = SO4T - DELTA
        IF (MIRAB .LT. TENAR) THEN

```

```

        X(41) = DELTA
        X(42) = 0
    ELSE
        X(42) = DELTA
        X(41) = 0
    END IF
1180 RETURN
END

```

C-----

C-----

SUBROUTINE MGSO4

C Subroutine for MgSO4 equilibrium.

```

REAL IX,I,K
CHARACTER*50 TITLE
DOUBLE PRECISION A,B,C,XA,MGT,SO4T,KPRIME,KIONPR,X,DELTA

COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
IX(30),TITLE,ITER,AX(60),TEMP,I,PHI

DELTA=0
MGT = X(4)+X(43)+X(44)+X(23)
SO4T = X(12)+X(43)+X(44)+X(23)
HEXAHY = K(43)/AH2O**6
EPSOM = K(44) / AH2O**7
IF (HEXAHY .LT. EPSOM) THEN
    KPRIME = K(43)/(GX(4)*GX(12)*AH2O**6)
ELSE
    KPRIME = K(44)/(GX(4)*GX(12)*AH2O**7)
END IF
IF (MGT*SO4T-KPRIME .GT. 0) GOTO 1200
1190 KIONPR = K(23)/(GX(4)*GX(12))
A = 1
B = -1*(MGT+SO4T+KIONPR)
C = MGT*SO4T
XA = (-B-SQRT(B**2-4*A*C))/(2*A)
X(4) = MGT - XA
X(12) = SO4T - XA
X(23) = XA
X(43) = 0
X(44) = 0
GOTO 1230
1200 DECREM = MIN(MGT,SO4T)
DO 1220 KA = 2, 5
1210     DELTA = DELTA+DECREM*(DEL(KA)-1)
        IF (DELTA .LT. 0) GOTO 1190
        KIONPR = K(23)/(GX(4)*GX(12))
        X(23) = KPRIME/KIONPR
        X(4) = MGT - DELTA - X(23)
        X(12) = SO4T - DELTA - X(23)
        IF (KA .EQ. 5) CALL PITZER
        HEXAHY = K(43)/AH2O**6
        EPSOM = K(44)/AH2O**7
        IF (HEXAHY .LT. EPSOM) THEN
            KPRIME = K(43)/(GX(4)*GX(12)*AH2O**6)
        ELSE
            KPRIME = K(44)/(GX(4)*GX(12)*AH2O**7)
        END IF

```

```

        IF (X(4)*X(12)-KPRIME .GT. 0) GOTO 1210
        DELTA = DELTA-DECREM*(DEL(KA)-1)
1220 CONTINUE
        X(23) = KPRIME/KIONPR
        X(4) = MGT - DELTA - X(23)
        X(12) = SO4T - DELTA - X(23)
        IF (HEXAHY .LT. EPSOM) THEN
            X(43) = DELTA
            X(44) = 0
        ELSE
            X(44) = DELTA
            X(43) = 0
        END IF
1230 RETURN
END

```

C-----
C-----

SUBROUTINE PICROMERITE

C Subroutine for K2Mg(SO4)2.6H2O equilibrium.

```

REAL IX,I,K
CHARACTER*50 TITLE
DOUBLE PRECISION A,B,C,XA,MGT,KT,SO4T,KPRIME,KIONPR,X,DELTA

COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
           XIX(30),TITLE,ITER,AX(60),TEMP,I,PHI

PHASE = PHASE + 1
DELTA=0
KPRIME = K(46)/(GX(4)*GX(2)**2*GX(12)**2*AH2O**6)
MGT = X(4) + X(46) + X(23)
KT = X(2) + 2*X(46)
SO4T = X(12) + 2*X(46) + X(23)
IF (MGT*KT**2*SO4T**2-KPRIME .GT. 0) GOTO 1250
1240 KIONPR = K(23)/(GX(4)*GX(12))
A = 1
B = -1*(MGT+SO4T+KIONPR)
C = MGT*SO4T
XA = (-B-SQRT(B**2-4*A*C))/(2*A)
X(4) = MGT - XA
X(12) = SO4T - XA
X(23) = XA
X(2) = KT
X(46) = 0
GOTO 1280
1250 DECREM = MIN(MGT,SO4T,KT)
DO 1270 KA = 2, 5
1260     DELTA = DELTA+DECREM*(DEL(KA)-1)
        IF (DELTA .LT. 0) GOTO 1240
        KIONPR = K(23)/(GX(4)*GX(12))
        X(23) = KPRIME/KIONPR
        X(4) = MGT - DELTA - X(23)
        X(12) = SO4T - 2*DELTA - X(23)
        X(2) = KT - 2*DELTA
        IF (KA .EQ. 5) CALL PITZER
        KPRIME = K(46)/(GX(4)*GX(2)**2*GX(12)**2*AH2O**6)
        IF (X(4)*X(2)**2*X(12)**2-KPRIME .GT. 0) GOTO 1260
        DELTA = DELTA-DECREM*(DEL(KA)-1)
1270 CONTINUE

```

```

X(23) = KPRIME/KIONPR
X(4) = MGT - DELTA - X(23)
X(12) = SO4T - 2*DELTA - X(23)
X(2) = KT - 2*DELTA
X(46) = DELTA
1280 RETURN
END

```

C-----
C-----

SUBROUTINE CASO4

C Subroutine for CaSO4 equilibrium.

```

REAL IX,I,K
CHARACTER*50 TITLE
DOUBLE PRECISION A,B,C,XA,CAT,SO4T,KIONPR,IXA,X

```

```

COMMON /B/ X(60),GX(60),K(20:60),AH2O,DEL(5),
XIX(30),TITLE,ITER,AX(60),TEMP,I,PHI

```

```

1290 CALL PITZER
KIONPR = K(22)/(GX(3)*GX(12))
CAT = X(3) + X(22)
SO4T = X(12) + X(22)
IXA = X(22)
A = 1
B = -1*(CAT + SO4T + KIONPR)
C = CAT*SO4T
XA = (-B-SQRT(B**2-4*A*C))/(2*A)
X(3) = CAT - XA
X(12) = SO4T - XA
X(22) = XA
IF ((ABS(IXA-XA)/XA)*100 .GT. 1) GOTO 1290
RETURN
END

```

C-----

REPORT DOCUMENTATION PAGE

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13. ABSTRACT (Maximum 200 words) This report documents a FORTRAN version of the Spencer-Møller-Weare chemical thermodynamic model for aqueous electrolyte solutions at subzero temperatures (FREZCHEM). FREZCHEM is structured to predict the chemical composition and unfrozen water of aqueous solutions between -60 °C and +25 °C at atmospheric pressure (0.101325 MPa). FREZCHEM includes two reaction pathways: 1) freezing at variable temperature and fixed total water and 2) evaporation at variable water and fixed temperature. Activity coefficients and the activity of water are calculated using the Pitzer equations, which are valid to high solution ionic strengths ($\approx 20 \text{ mol kg}^{-1}$). Fifteen chloride and sulfate salts of sodium, potassium, calcium, and magnesium are included in the model. Predicted and experimental measurements of solute molalities and the unfrozen water fraction during seawater freezing are in good agreement. At -50 °C, 0.3% of seawater remains unfrozen with 99.7% of Na and 95.5 % of Cl having precipitated into one of four salts. FREZCHEM should find many applications in physicochemical studies of aqueous solutions and freezing.			
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